

Controllable synthesis of gold nanoparticles in aqueous solution by microwave assisted flow chemistry

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Abstract

The development of energy efficient, reproducible and high throughput approaches to gold nanoparticles (Au-NPs) synthesis have gained increasing attention over the past decades due to their application in biomedicine, sensor and catalysis. In this work, single mode microwave irradiation is for the first time combined with microflow chemistry to fabricate Au-NPs continuously and reproducibly with controllable size in an aqueous solution. The major experimental parameters including microwave power, citrate-to-gold molar ratio ($[Cit]/[Au]$) and reaction residence time have been investigated systematically. As indicated by TEM, the mean particle width of the synthesised Au-NPs is between 4 and 15 nm with mean aspect ratio between ~ 1.4 and 2.2 after only 90 s microwave irradiation. Furthermore the Au particle morphology can be manipulated from nanowires to nanoparticles by adjusting the $[Cit]/[Au]$ ratio. HRTEM analysis of the produced Au-NPs and UV-vis spectroscopy suggest a correlation between the red-shifted surface plasmon resonance peak and the particle aspect ratio that is probably dependent on the creation of particle-particle junctions, which alter both the size and the aspect ratio of the Au-NPs. The synthesis further demonstrated high reproducibility.

Keywords: microwave, flow chemistry, gold nanoparticles, particle size, aspect ratio

Introduction

Gold nanoparticles (Au-NPs) have attracted increasing attention over the last decades due to their unique surface plasmon resonance properties, making them suitable in many areas such as healthcare and sensing applications.^{1,2} Reduction of HAuCl_4 in boiling water by sodium citrate (Na_3Cit) which acts as both a reducing agent and a stabilizer is widely used for Au NPs synthesis (Turkevich method).³ For this purpose, a conventional heated batch reactor was often applied.⁴⁻⁹ In order for high yield synthesis of Au-NPs with well-defined shape and size, the Turkevich method has been extensively modified such as by changing the addition order of the reactants, using different reducing agents, adjusting the ratio of HAuCl_4 to sodium citrate and pH values as well as adding new stabilisers.¹⁰⁻¹⁸ Replacement of water by its expensive isotope, deuterium oxide, was shown to produce smaller sizes of Au NPs (5.3 ± 1.1 nm), compared with the particles (9.0 ± 1.2 nm) produced in pure water which was attributed to faster reduction of precursor.¹⁹ An increase in reaction rate was observed by reversing the order of the reactants' addition (i.e. HAuCl_4 to Na_3Cit solution) that probably accelerated the thermal decomposition or oxidation of sodium citrate, resulting in the formation of Au-NPs in the 5–10 nm size range.^{11, 12} Recently, synthesis of sub-10 nm citrate-stabilized Au NPs has been achieved following a kinetically controlled seeded-growth strategy, combining tannic acid with sodium citrate during nucleation.²⁰ However, one of the major challenges in conventional batch systems has always been the non-uniform distribution of heat that results in temperature fluctuation, causing many variances in Au-NP synthesis processes, particularly in the nucleation and growth steps,²¹ resulting into wide particle size distribution. Furthermore the batch to batch difference also raise concerns about reproducibility and scaling-up issues.²¹

The growth mechanism of the Turkevich method in a batch reactor and the influences of reagent ratio, reaction temperature and pH of the medium on synthesis of Au-NPs have recently been studied

comprehensively in order to further control the nucleation and particle growth and then final particle size and size distribution.²² In this study, it was shown that the Turkevich synthesis followed a four-step seed-mediated growth mechanism in which no large particles (>25 nm) were formed during the synthesis, confirmed by *in-situ* ultrasmall angle X-ray scattering.²² In contrast to other proposed mechanisms,²³ their observations excluded the formation of any large aggregates (*e.g.* >25 nm) throughout the Turkevich synthesis.

As an emerging heating technology, microwave heating (MH) has been shown to dramatically reduce the reaction processing time and provide homogenous bulk heating to synthesise various distinctive nanocrystals with defined structures. Similar to the synthesis of other metal nanoparticles,^{7, 8} MH was employed in a batch reactor to prepare Au-NPs with various shapes including spherical, polygonal, rod, and triangular/prisms following the Turkevich or other synthesis routes using different reagents (*e.g.* 2,7-dihydroxy naphthalene as the reducing agent and cetyltrimethyl ammonium bromide as the stabilizer).^{16, 24-27} These studies generally used domestic multimode microwave systems (600-1000 W) and their results revealed that microwave assisted fast temperature ramping was highly effective to synthesise Au-NPs (*e.g.* 20.01 (\pm 2.05) nm or 12.04 (\pm 1.35) nm). Moreover, the results showed that the final products could vary in size and morphology depending on the type of reactor used, microwave absorption ability, reaction time *etc.* However, there is an obvious drawback in these reports that the particle size distribution and yield are not satisfactory due to the limited penetration depth of microwave irradiation (~1 cm in aqueous solution) and large reactor used (normally larger than 2cm in diameter), which also affects scalability of the synthetic processes.^{7, 8, 28}

It is well known that single mode microwave irradiation can transfer energy to the reaction medium much more efficiently and faster compared to multimode counterparts (*e.g.* domestic microwave systems).^{29, 30} Faster temperature ramping rates are likely to produce more seeds by promoting

nucleation instead of particle growth, therefore smaller Au-NPs could be potentially achieved with a high yield by a single mode microwave irradiation even at very low microwave power. Furthermore, micro channel reactors were reported to have a better heat transfer rates compared with large diameter reactors due to the increased surface area-to-volume ratio.³¹ As such, one can anticipate that single mode microwave heating coupled with a micro-flow system which has a diameter smaller than the microwave penetration depth (< 1 cm) should be potent to prepare Au-NPs with smaller particle size distribution and high yield. Given such an attractive approach, to the best of our knowledge, there is not a continuous microflow system driven by microwave irradiation reported to synthesise Au-NPs. Herein, a novel single mode microwave irradiation promoted microflow chemistry (SMMM) is presented to synthesise Au-NPs continuously with high reproducibility. Furthermore, the major variables e.g. microwave power, citrate-to-gold precursor molar ratio, residence time are comprehensively investigated and finally the SMMM is also compared with conventional flow system.

Experimental

The SMMM used consisted of a flow reactor (6 mL) made of 1/8" OD (~0.32 cm) Teflon tubing, a commercially available single-mode microwave system (Discover, CEM Microwave Technology Ltd), a syringe pump (KD Scientific), a thermocouple temperature probe positioned 5 cm away from the exit of microwave heating zone, a sample and waste collection unit with backpressure regulator (20 psi, IDEX Health and Science) to control reactor pressure and an ice-bath placed between the end of the flow reactor and sample collection unit.

Two precursors hydrogen tetrachloroaurate (III) hydrate ($\text{HAuCl}_4 \cdot 3\text{H}_2\text{O}$, 99.9%) and trisodium citrate dehydrate ($\text{Na}_3\text{C}_6\text{H}_5\text{O}_7$, 99%) were purchased from Sigma-Aldrich and used as received. In a typical protocol, 2 mL of freshly prepared solution of HAuCl_4 (0.005 M) was added to a graduated cylinder and

diluted to 25 mL using deionised water. The required amount of Na₃Cit (0.05M) to prepare the targeted citrate-to-gold molar ratio was then added to HAuCl₄ aqueous solution, and the mixture of HAuCl₄ and Na₃Cit was diluted to 40 mL using deionised water to get 0.25 mM of HAuCl₄ concentration. This mixture of HAuCl₄ and Na₃Cit was immediately transferred into a 60 mL syringe. The mixture of HAuCl₄ and Na₃Cit then flowed through the reactor under a constant pressure. Residence times of approximately 0.6, 0.85 and 1.5 min that correspond to flow rates of 10, 7 and 4 mL/min were used, respectively. The hot fluid containing formed Au-NPs was immediately cooled by an ice-bath to quench any further reactions that may affect the homogeneity of Au-NPs product.

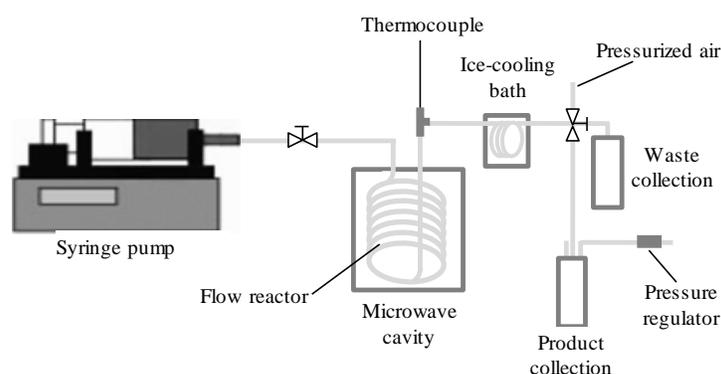
Sample Characterization

UV-vis analysis of Au-NPs was performed immediately after material synthesis and concurrently the samples for TEM analysis were prepared. Transmission electron microscopy (TEM) and high resolution transmission electron microscopy (HRTEM) were performed using Jeol JEM-1010 and JEOL-2010F coupled with an EDS detector (Oxford Instruments), respectively. TEM samples were prepared by dropping fresh solution of synthesized Au-NPs (20 µL) on carbon coated TEM grids unless stated otherwise. Ultraviolet-visible (UV-Vis) absorption was recorded by Shimadzu UV-2550 UV-vis spectrophotometer. A typical Au-NP solution in de-ionized water (~3 mL) was placed in a cell, and spectral analysis was performed in the range of 200-900 nm. Size of Au-NPs was measured via ImageJ software, from their TEM images. The particle dimensions were measured in-plane, assuming that the width of the particles are comparable to their height. In case of non-spherical particle morphologies (*e.g.* boomerang and T-shape fused particles) observed, the longest particle dimensions either the length or width of the particle was measured to calculate particle aspect ratio. Descriptive statistical analysis was performed for individual data sets (Weight method: Direct weight, Variance divisor of moment: DF,

Interpolation of Quantiles: Empirical distribution with averaging) using Origin 9.1 software to find out mean particle size and aspect ratio.

Results and Discussion

The SMMM was built up as depicted in **Scheme 1** and used for continuous Au-NPs synthesis. Direct measurement of the temperature of hot fluid inside the micro-reactor is very challenging due to the size and configuration of the micro-reactor and small microwave cavity. Therefore, temperature of reaction fluid was recorded by using a thermocouple temperature probe placed inside T-type PTFE tubing and positioned 5 cm away from the exit of microwave irradiated zone. The temperature of the fluid in microwave zone is expected to be higher than that of the temperature recorded outside the microwave zone. The likely temperature in the microwave zone (just at the exit of the microwave irradiated zone) was estimated by considering forced convection heat transfer between the fluid inside the PTFE tubing and the outside surrounding air (see **ESI** for calculations). Heat loss due to radiation was neglected due to relatively low temperature of the fluid (~ 120 °C). The calculated values of the fluid temperature at different fluid flow rates are listed in ESI Table S1.



Scheme 1 Schematic representation of the combined single mode microwave irradiation and microflow technology to synthesize Au-NPs.

Influence of microwave irradiation intensity

Formation of citrate stabilized Au-NPs was previously shown to follow a mechanism, including nucleation, aggregation, and spherical particle growth via Oswald ripening.^{14, 22, 32, 33} For microwave assisted Au-NP synthesis, it is widely accepted that microwave irradiation accelerates the nucleation of particles due to its selective and fast heating characteristics.^{24-27, 34} We have also recently found that a SMMM system has similar advantages including fast temperature ramping rate and selective heating of iron oxide for its self-assembly.³⁰ Change in temperature ramp rates due to varied microwave powers is therefore expected to influence the nucleation rate of nanoparticles. Previous studies showed that Au-NPs with uniform size $\sim 12.04 \pm 1.35$ nm could be produced in >5 min after reaching 95 °C reaction temperature using microwave heating with fast ramping rate at >50 °C min^{-1} and maintaining the pH value as >6.2 by adjusting the $\text{Na}_3\text{Cit}/\text{HAuCl}_4$ molar ratio ~ 5.68 .²⁷ Based on these results, in order to examine the effect of microwave power, Au-NPs were synthesized at microwave powers of 20, 25, 30, 36, 40 and 45 W, while maintaining citrate-to-gold molar ratio (6/1), flow rate (4 mL min^{-1}) and back pressure (20 psi) constant. Temperature profiles of the aqueous solution were monitored for 5 min by measuring the temperature of the flowing reaction fluid outside the microwave zone at 30 second interval (see **ESI Figure S1**). Note that the volume of the reactor used was approximately 6 mL, and the residence time of reaction fluid in the reactor was very short (e.g. 90 s at 4 mL/min flow rate). In agreement with previous literature,^{27,29} temperature in SMMM system reached a certain value faster when higher microwave power was applied, compared to low microwave power (see **ESI Figure S1**). The temperature of the reaction fluid measured at the exit of microwave zone was ~ 117 °C in approximately 2 min flow time and thereafter remained constant when microwave power was larger than 30 W (e.g. 36, 40 and 45 W). On the other hand, temperature of the flowing reaction fluid did not stabilize when 20, 25 or 30 W microwave power was applied. It showed a slightly increasing

temperature trend between 2 min and 5 min (see **ESI Figure S1**). Clearly visible, coloured aqueous solutions of Au-NPs with increasing colour intensity were obtained in the case of 25, 30, 36, 40 and 45 W applied microwave power while the solution for 20 W applied microwave power was almost colourless, probably due to inadequate heating at low microwave power. Therefore in following studies 20 W microwave power was not considered and the sample prepared at this condition was not analysed. All coloured solutions were initially characterised by UV-vis absorption spectroscopy. Characteristic surface plasmon resonance (SPR) peaks of the obtained Au-NPs confirm continuous Au-NPs formation under 25, 30, 36, 40 and 45 W applied microwave power (**Figure 1a**). Au-NPs synthesized at microwave power from 25 W to 45W display UV-vis absorption maxima positioned at approximately 560, 547, 542, 545 and 543 nm, respectively (**Figure 1a and c**). Furthermore, a weak shoulder is observed at *ca.*750 nm for Au-NPs synthesized under 25 W, that may probably be attributed to the presence of high aspect ratio (>2) Au-NPs observed in TEM images (**ESI Figure S2**).^{34, 35} Typical TEM image of the Au-NPs synthesized under constant citrate-to-gold molar ratio of 6/1 and 4 mL min⁻¹ flow rate at 36 W is shown in **Figure 1b**. The high magnification HR-TEM image (inset of **Figure 1b**) proves Au-NPs (average aspect ratio of 1.37) with lattice spacing consistent with Au⁰ (111) have been synthesised by the system.³⁶

Careful inspection of the absorption spectra of Au-NPs reveals that the maximum absorption (λ_{max}) of Au-NPs synthesized at high microwave power (e.g. 45W) is blue-shifted compared to low microwave power (e.g. 25W) as shown in Figure 1c. For spherical Au-NPs, a blue-shift in λ_{max} is usually attributed to the formation of Au-NPs smaller in size.^{27, 37} However, UV-vis spectroscopy alone is an inadequate technique for particle size and morphology analysis since the SPR of Au-NPs is highly morphology dependent, thus the use of alternative techniques such as electron microscopy is necessary to elucidate

the reason of the observed λ_{\max} shift.^{9, 26, 27, 34, 35, 38} Thus transmission electron microscopy (TEM) has been extensively used to observe the mean particle size and size distribution of Au-NPs.^{13, 15, 35}

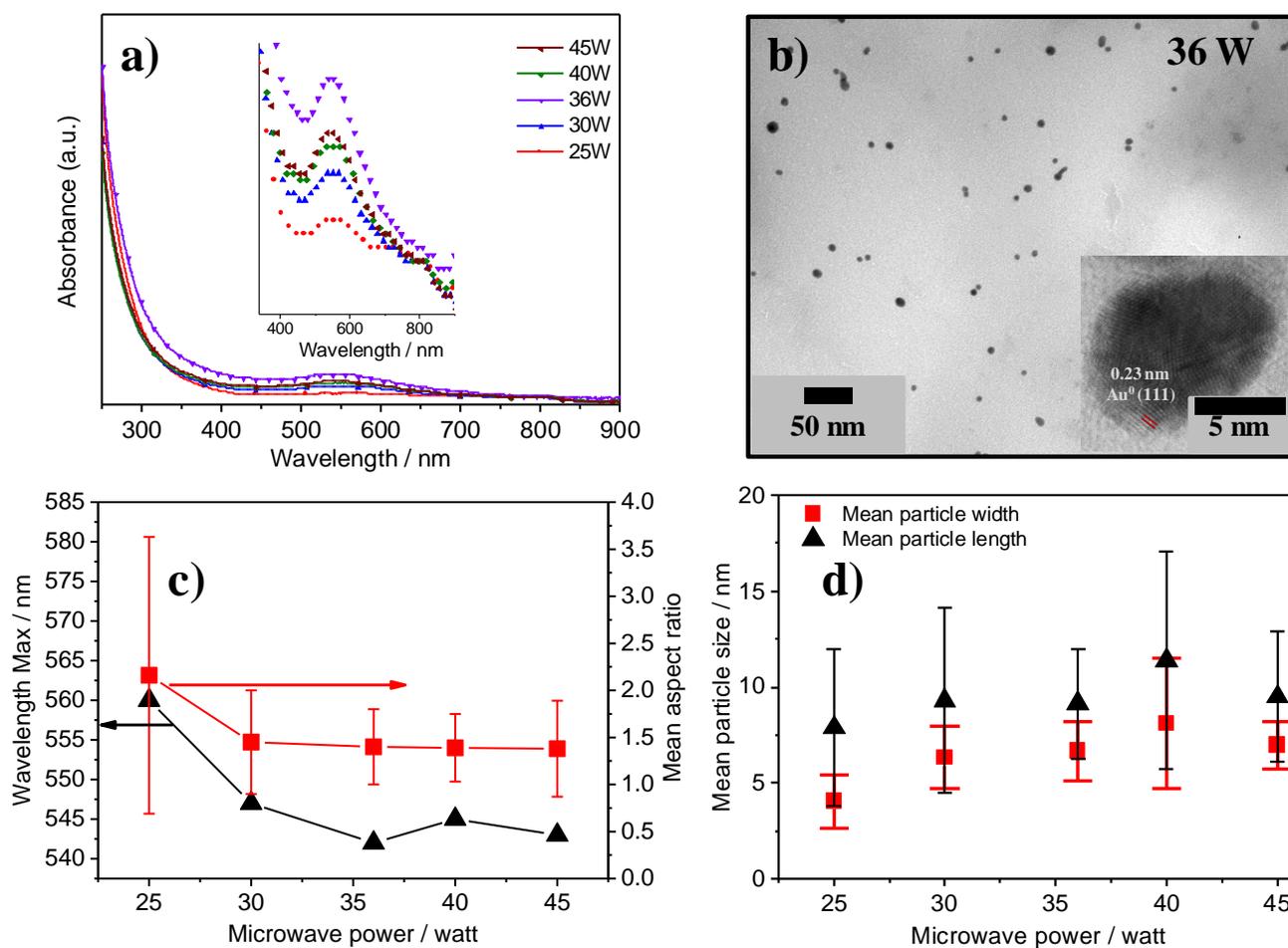


Figure 1 a) UV-vis spectra of the Au-NPs synthesised under the conditions given above. Inset: Enlarged UV-vis spectra of the Au-NPs synthesised. **b)** TEM images of the Au-NPs synthesized under constant citrate-to-gold molar ratio of 6/1 and 4 mL min⁻¹ flow rate by applying microwave power of 36 W. Inset: High magnification HR-TEM image shows a selected non-spherical Au-NP (aspect ratio of 1.37) with lattice spacing (0.23 nm) consistent with Au⁰ (111). **c)** Graph shows the wavelength of maximum

absorption (λ_{\max}), the mean aspect ratio and the standard deviations in aspect ratio of the Au-NPs synthesized by applying microwave power of 25, 30, 36, 40 and 45 W. **d)** Graph shows the mean particle length/width and the standard deviations in size of the Au-NPs synthesized by applying microwave power of 25, 30, 36, 40 and 45 W. Solid curves are intended to guide the eye.

TEM analysis of Au-NPs synthesized by the SMMM system shows the presence of a mixture of spherical and high-aspect ratio non-spherical Au-NPs, rather than commonly observed spherical Au-NPs formation via Turkevich method^{12, 22, 39-41} (see **ESI Figure S2** for TEM images). Further particle size and morphology were analysed by using corresponding TEM of the individual samples obtained from at least six different regions of TEM grid. The mean particle length, the mean particle width and the mean aspect ratio of Au-NPs and their standard deviations were then statistically obtained (**Figure 1 c and d**), which is also summarized in ESI Table S2.

The maximum absorption (λ_{\max}) is blue shifted when 25W to 36W of microwave power is applied. Thereafter the changes become small when microwave power further increases, while there is slightly red shift when 40 W of microwave power was applied (**Figure 1c**). Interestingly, the change of λ_{\max} is relatively well correlated to the change of aspect ratio of the prepared Au-NPs at 25, 30, 36 and 45 W (**Figure 1 b and c**). Furthermore, the ratio of sphere-like particles to elongated particles were counted and shown in **ESI Figure S3a**. Notably there is lower percentage of spherical Au-NPs (40.5%) prepared at 40 W which was derived using the amount of Au-NPs with aspect ratio smaller than ~ 1.2 , compared to the Au-NPs prepared at 36 (41.8%) and 45 W (47.7%). In other words, there are more elongated particles when 40W was applied, which causes the small red shift in the λ_{\max} and also explain the larger mean particle length/width.

Citrate stabilized spherical Au-NPs using conventional methods were previously reported by a few groups and, it was shown that the λ_{\max} values of Au-NPs with particle size of 5-20 nm were in the range of 515-530 nm.^{15, 39} In addition, microwave assisted batch synthesis of spherical Au-NPs with size distribution of 12.04 ± 1.35 nm was reported, showing that smaller particles were produced at high microwave powers, attributed to higher supersaturation ratio of Au atoms in the solution at faster temperature ramp rates,²⁷ which might be the case when 45W microwave power was applied in this study. For spherical Au-NPs (aspect ratio ~ 1), observed red shift in SPR peak position was generally attributed to change in the particle size or the variations in the refractive index of the surrounding medium of the nanoparticles.⁴² In contrast to previously reported citrate stabilized spherical Au-NPs (aspect ratio ~ 1) with λ_{\max} at *ca.* 515-530 nm,^{15, 27, 39} the Au-NPs synthesized under 36, 40 and 45 W of microwave powers in a SMMM system show λ_{\max} at *ca.* 542-545 nm although their size is below 10 nm (**ESI Table S2**). Therefore large λ_{\max} for Au-NPs produced in the SMMM system is attributed to the relatively large aspect ratios in the current study (**Figure 1c** and **ESI Table S2**).

Influence of citrate-to-gold molar ratio ([Cit]/[Au])

Given that the relatively high intensity UV-vis absorption of Au-NPs achieved at microwave power of 36 W indicated by its UV-vis intensity, Au-NPs were further synthesized continuously by varying [Cit]/[Au] in the range 1-60 at fixed microwave power of 36 W and 4 mL min⁻¹ flow rate. The temperature just outside the microwave zone increased to ~ 118 - 119 °C in approximately 2 min fluid flow for all [Cit]/[Au] ratios used (**ESI Figure S4**). Aqueous solutions of Au-NPs were collected after the outside temperature remained at ~ 118 - 119 °C (after 2 min). Au-nanowire-like structures with polygonal heads (width \times length: $\sim 10 \times \geq 200$ nm) were obtained in a short time (2 min) when [Cit]/[Au] of 1/1 is used (**Figure 2**), probably due to preferential adsorption of AuCl₄⁻ on the surface of the initially

formed thermodynamically unstable gold nanoparticles. This process is favourable when the stabiliser citrate ions are not enough to fully cover their surface.

³³ Similar gold nanowire formation at 80 °C in 30 min was observed in a batch reactor by citrate reduction of AuCl_4^- with a low concentration of citrate³³ although it has recently been reported that no particles longer than 25 nm have been observed during Au-NPs synthesis using the Turkevich method.²²

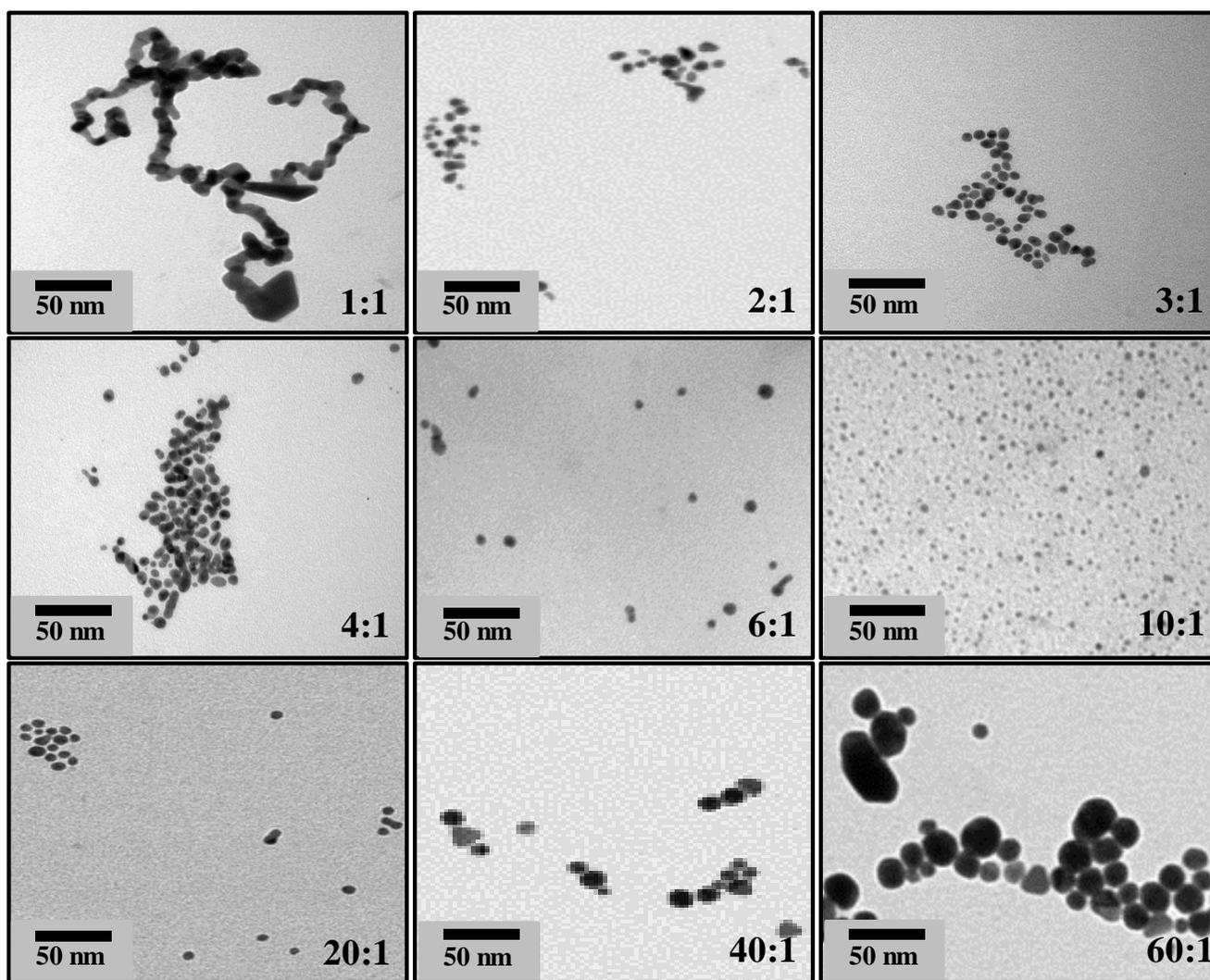


Figure 2 Typical TEM images of the Au-NPs synthesized under constant microwave power of 36 W, 4 mL min⁻¹ flow rate by using citrate-to-gold molar ratios of 1, 2, 3, 4, 6, 10, 20, 40 and 60.

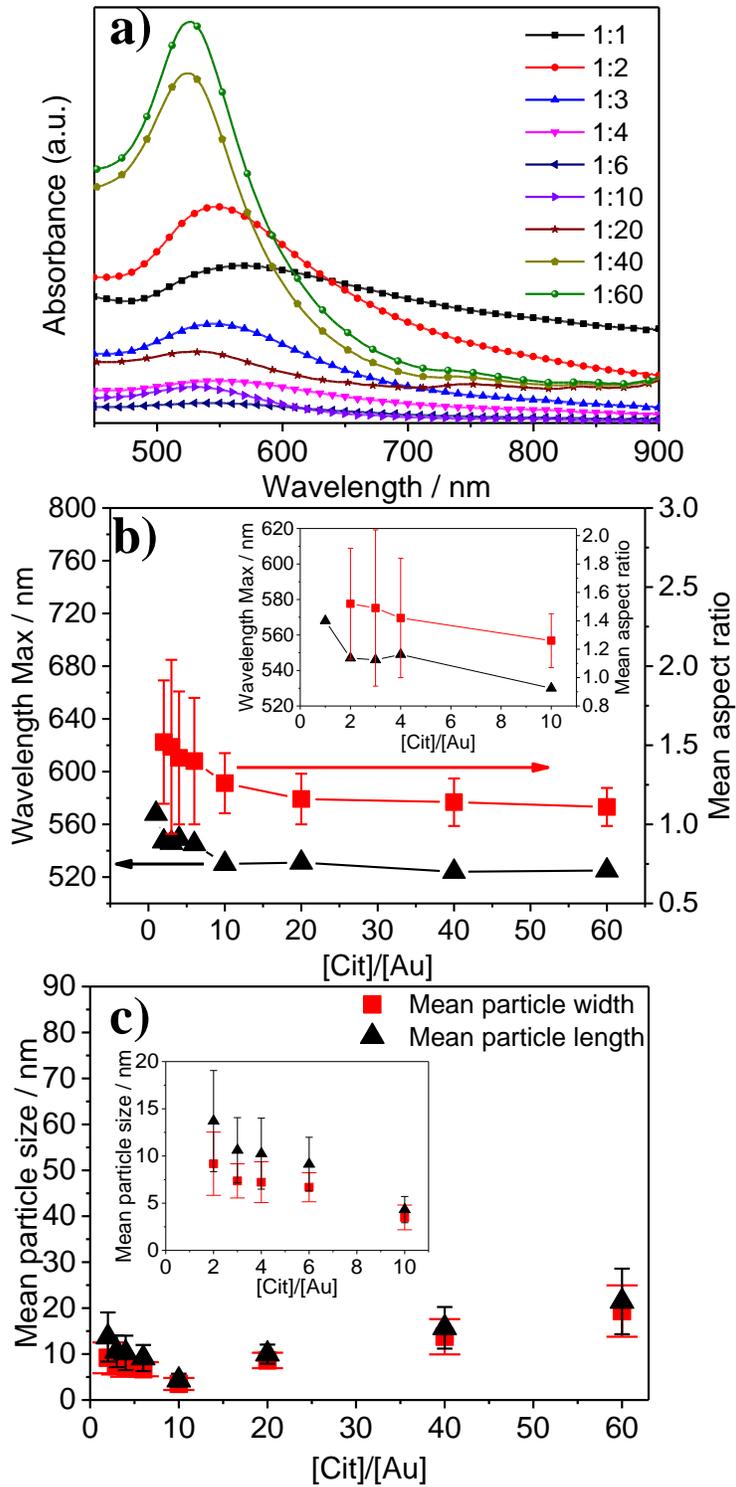


Figure 3 a) UV-vis spectra of the Au-NPs synthesized under constant microwave power of 36 W, 4 mL min⁻¹ flow rate and 20 psi back pressure by using citrate-to-gold molar ratios of 1, 2, 3, 4, 6, 10, 20, 40 and 60. **b)** Graph shows the wavelength of maximum absorption (λ_{\max}) and the mean aspect ratio and the standard deviations in aspect ratio of the Au-NPs synthesized by using [Cit]/[Au] of 2, 3, 4, 6, 10, 20, 40 and 60. Note that, for [Cit]/[Au] of 1/1, only λ_{\max} absorption was presented since aspect ratio of gold-nanowires was not calculated. Inset shows the region from [Cit]/[Au] of 2 to 10. **c)** Graph shows the mean particle length/width and the standard deviations in size of the Au-NPs synthesized by using [Cit]/[Au] of 2, 3, 4, 6, 10, 20, 40 and 60. Note that there is no data point for [Cit]/[Au] of 1/1. Inset shows the region from [Cit]/[Au] of 2 to 10. Solid curves are intended to guide the eye.

After the TEM observation, UV-vis analysis of produced Au-NPs were carried out. It shows λ_{\max} at *ca.* 547, 546, 549, 545, 530, 531, 524 and 525 nm for [Cit]/[Au] ratios of 2, 3, 4, 6, 10, 20, 40 and 60, respectively (**Figure 3 a and b**). Consistent with the literature,^{35, 43} as-synthesized Au-nanowires using [Cit]/[Au] of 1/1 show a broad UV-vis absorption band (λ_{\max} ~568 nm) with a long tail in near IR region, characteristic of Au-nanowires (**Figure 3 a and b**). It is known that high aspect ratio elongated Au-NPs possess both SPR band positioned at *ca.* 520 nm (so-called transverse band), characteristic of spherical Au-NPs, and another absorption band, so-called longitudinal band, at relatively longer wavelengths owing to the plasmon oscillation of electrons along the long axis of the elongated Au-NPs.^{35, 42} These bands are clearly separated from each other for the high aspect ratio well-defined Au-NPs. In addition, they may significantly get closer in some circumstances and may appear as a broadened single band for the low aspect ratio (or spherical) Au-NPs.³⁵ However, in our study formed gold nanowires show a board peak observed from 500 nm to 900 nm with a significantly flattened absorption pattern after 800

nm. Consistent with the previous report, this broad profile may be attributed to the superposition of the longitudinal resonance of gold nanowires with varying lengths at this wavelength range.³³

The statistically derived mean particle size of the produced Au-NPs decreases when [Cit]/[Au] ratios increase from 2/1 to 10/1 and reaches its minimum at the ratio of 10/1 (**Figure 3c**). After that, the mean particle size of the produced Au-NPs shows an increasing trend. Thus the smallest (mean particle length=4.3 nm; mean particle width=3.5 nm) and the largest (mean particle length=21.5 nm; mean particle width=19.4 nm) Au-NPs are produced when [Cit]/[Au]s of 10/1 and 60/1 are used, respectively (see **ESI Table S3**), which is agreement with Au NPs prepared in a batch reactor by conventional heating^{14, 44} Furthermore, a clear dependency of particle size of produced Au-NPs on citrate ratio was observed at constant H₂AuCl₄ precursor concentration (0.25 mM) in the SMMM system. In other words, the concentration of citrate has non-monotonic effect on the final size of the produced Au-NPs, as reported by Kumar *et al.*⁴⁵ and Polte *et al.*²². It is proposed that the change in particle size with citrate concentration is probably due to the capping effect of the citrate ions.²⁷ More citrate ions, better capping effect, leading to smaller particles. Further increasing citrate ions concentration, reduction of gold precursor is dominated which potentially may accelerate particle growth. The detailed reason is currently underway.

One can observe there is a correlation between the particle size and λ_{\max} absorption (**Figure 3b** and **c**). The mean aspect ratio of the produced Au-NPs shows a decreasing trend when [Cit]/[Au] ratios increase from 2/1 to 20/1. Thereafter, it remains relatively unchanged. Again a correlation between the particle aspect ratio and λ_{\max} absorption is observed, the higher the aspect ratio the larger the λ_{\max} , (**Figure 3b**). In particular, this correlation is observed when two materials prepared under different condition but with similar mean particle lengths. For instance, the Au NPs prepared at [Cit]/[Au] ratios of 4/1 and 20/1 show mean particle length of 10.3 and 9.9 nm. (see **ESI Table S3**). Although the mean particle lengths

are comparable, their λ_{\max} absorptions are observed at 549 and 531 nm, respectively, corresponding to mean aspect ratio of 1.42 and 1.16 (**Figure 3b**). Therefore we would like to propose that the aspect ratio dominates to some extent to the λ_{\max} absorption of the Au-NPs synthesized by the SMMM system. A previous study showed the effect of the Au-NPs geometry on the localized surface plasmons (LSPs) theoretically and experimentally. By comparing large oblate, prolate and ellipsoid shaped Au-NPs (>50 nm)⁴⁶ the study indicated that the LSPs were highly sensitive to change in the dimensions of ellipsoid particles, suggesting a correlation between the aspect ratio and the λ_{\max} absorption of the Au-NPs. It was also reported that the overlapped particles (or junctions), so-called hot-spots occurring between a pair of adjacent nanoparticles, could result in highly intense and localized electromagnetic fields upon excitation by incident light of the appropriate polarization. The region of the overlap between particles or the distance between two adjacent nanoparticles was responsible for the shift of the SPR peak.⁴⁷⁻⁵¹ In agreement with these studies, close inspections on HRTEM images of the Au-NPs produced by the SMMM system clearly reveal the formation of interconnected particles (see **ESI Figure 5**). The interconnection would alter the size of the resulting Au-NPs, which is probably responsible for increasing of the aspect ratio and the observed red-shift in the λ_{\max} absorption of the Au-NPs. It is also worth mentioning that the produced largest Au-NPs produced (21.5±7.1 nm/19.4±5.6 nm; length/width) prepared using [Cit]/[Au] ratios of 60/1 show relatively blue-shifted λ_{\max} absorption compared to Au-NPs produced using other [Cit]/[Au] ratios, again likely due to the amount of the highest percentage of Au-NPs with aspect ratio smaller than ~1.2, which is 90% (see **ESI Table S3** and **Figure S3b**).

Influence of flow rate

Furthermore we studied the influence of reactant flow rates on the produced Au-NPs when other reaction parameters (reactor volume 6 mL, microwave power 36 W and applied back pressure 20 psi)

were kept constant. The citrate-to-gold ratio of 60/1 was used since it produced Au-NPs with the highest UV-vis absorption intensity at 525 nm (see **Figure 3a**), and the reaction was carried out at flow rates of 4, 7 and 10 mL min⁻¹. 4 mL min⁻¹ flow rate was found as the reliable flow rate in the current configuration (see **ESI Figure S6a-d**). A clear red solution confirming Au-NP formation was obtained at flow rate of 4 mL min⁻¹, compared to 7 and 10 mL min⁻¹ which yielded a cloudy grey solution (see photographs in **Figure S6a-c**). TEM image of the red solution displays Au-NPs with mean length/width size of 21.5±7.1 nm/19.4±5.6 nm (**Figure S6a**). In contrast, TEM analysis of the solutions obtained at flow rate of 7 and 10 mL min⁻¹ shows a mixture of Au-NPs formed of spherical, high aspect ratio and rod-like particles (**Figure S6b** and **6c**). Electron microscopy results reveal that flow rate is critical to produce Au-NPs with well-defined morphology under given experimental conditions. Complementary UV-vis spectroscopy further supports the formation of Au-NPs at 4 mL min⁻¹ flow rate, showing a strong SPR band at 524 nm, compared to a weak one at 575 nm for 7 mL min⁻¹ flow rate (**Figure S6d**). However, no clearly visible SPR band is observed in the enlarged UV-vis spectrum (**Figure S6d**) for 10 mL min⁻¹ flow rate although TEM analysis showed some particles' formation (**Figure S6c**), probably due to very low concentration of Au-NP at such a short reaction time.

Reproducibility of synthesis

Reproducibility of the SMMM system was tested by repeating the fabrication of Au-NPs three times under microwave power of 36 W, 4 mL min⁻¹ flow rate and 20 psi back pressure by using citrate-to-gold molar ratios of 6/1. Au-NPs were characterized by both UV-vis spectroscopy and TEM. In all repeated experiments UV-vis analysis of produced particles shows absorption maxima at *ca.* 542 nm (**Figure 4a**) and, TEM analysis indicates Au-NPs with particle size distribution of ~10 nm (length) and ~7 nm (width) (**Figure 4b-c**), confirming that the novel synthesis method is reproducible and capable of producing Au-NPs with similar size distributions, which is extremely important for nanomaterial manufacturing

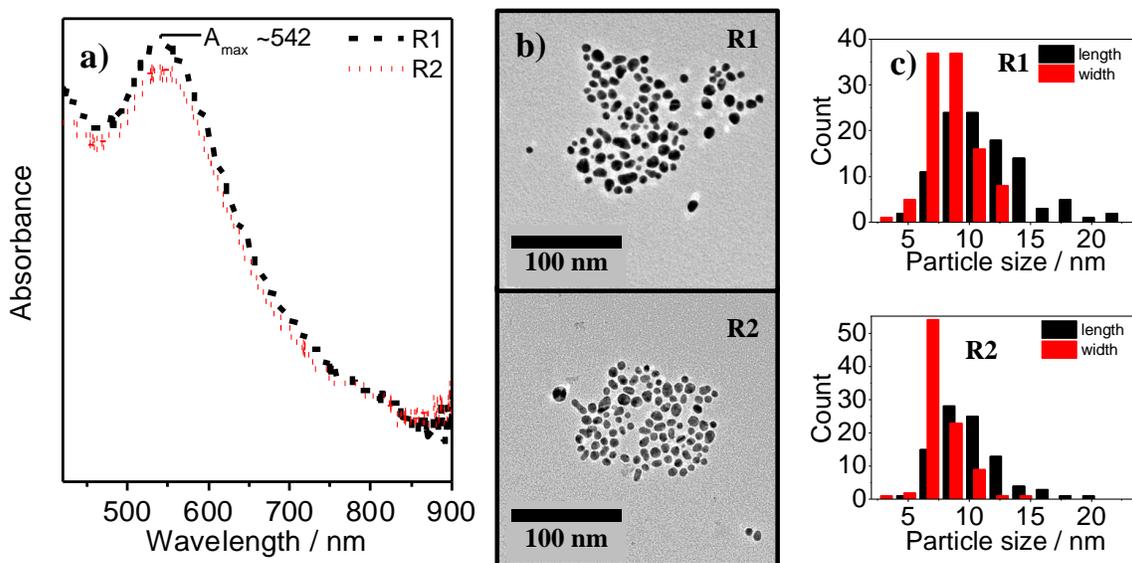


Figure 4 a) UV-vis spectra of the Au-NPs synthesized under constant microwave power of 36 W, 4 mL min⁻¹ flow rate by using citrate-to-gold molar ratios of 6/1. b) TEM images of Au-NPs produced by two different repeat experiments (R1 and R2) under the same experimental conditions and c) the particle size distribution of the produced Au-NPs.

On the other hand, Au-NPs prepared by conventional heating at 100 °C under identical experimental conditions (citrate-to-gold ratio of 6/1, 4 mL min⁻¹) show mean length/width size of 48.7 (\pm 18.6) nm/35.9 (\pm 10.1) nm, suggesting the superiority of microwave assisted fast nucleation in a flow system over conventional heating (Figure 5).

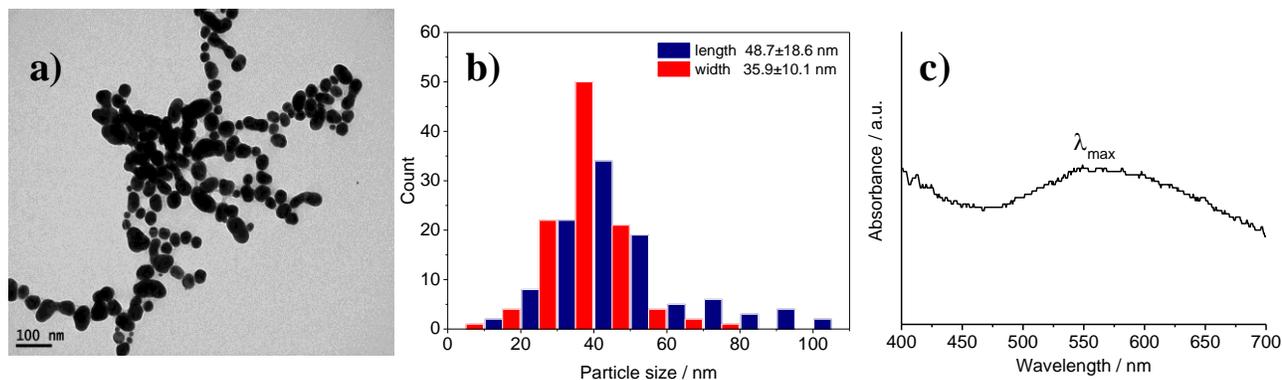


Figure 5 a) TEM images of the Au-NPs synthesized at constant citrate-to-gold molar ratio of 6/1, 4 mL min⁻¹ flow rate and 20 psi back pressure by *conventional heating* at 100 °C. b) The particle size distribution of the produced Au-NPs. c) UV-spectrum of the produced Au-NPs. λ_{\max} is observed at *ca.* 550 nm.

Conclusion

For the first time, a combining single mode microwave irradiation and microflow system (SMMM) has successfully been used to synthesize Au-NPs with controlled particle size (mean particle width \times length in the range of $\sim 4 \times 8$ nm and $\sim 8 \times 12$ nm) and mean aspect ratio (between ~ 1.4 and 2.2) with just 90 s reaction time. Au-nanowire-like structure with polygonal heads (width \times length: $\sim 10 \times \geq 200$ nm) is produced at the citrate-to-gold ratio of 1/1. Increasing the citrate-to-gold ratio from 2/1 to 10/1 leads to a decrease in particle size, however further increase from 10/1 to 60/1 results in an increase in particle size which is likely due to dual function of citrate ions. Interestingly, Au-NPs smaller than 15 nm with relatively higher mean aspect ratios ($> \sim 1.4$) show red-shifted SPR peaks (> 540 nm) compared to those (< 530 nm) previously reported with similar particle size of Au-NPs but mean aspect ratios closer to ~ 1 , suggesting the strong dependency of SPR peak on the aspect ratio of produced Au-NPs, which is further witnessed by the correlation of the maximum absorption (λ_{\max}) with aspect ratio achieved in three different conditions, changing microwave power, precursor ratios and flow rates. It is also proposed that the change in the particle aspect ratio is related to the formation of particle-particle junctions (observed by HRTEM), which is expected to have an influence on both the particle size and the aspect ratio. In addition, the novel synthesis approach has been proved reliable and reproducible.

In brief, the developed SMMM system combining the advantages of both single mode microwave heating and flow chemistry allows to control the reaction variables (*e.g.* heat, temperature, cooling, reactant concentration, time, reaction volume *etc.*) immediately and effectively during the course of the reaction, resulting into much smaller particle size and size distribution and at the same time benefiting from

fast and high throughput synthesis with good reproducibility between runs. It has also a strong potential to continuously fabricate metal and metal oxide nanoparticles in an economical manner.

Supporting Information

Fluid temperatures outside the microwave heating zone, tables of mean particle length and width and mean aspect ratio of the produced Au-NPs by using different microwave powers or citrate-to-gold molar ratios, TEM images of these particles synthesized at different microwave powers or flow rates, and UV-vis spectra of the Au-NPs synthesized at varying fluid flow rates.

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