

Corrigendum: Visible-Light-Driven Hydrogen Evolution Using Planarized Conjugated Polymer Photocatalysts . R. S. Sprick, B. Bonillo, R. Clowes, P. Guiglion, N. J. Brownbill, B. J. Slater, F. Blanc, M. A. Zwijnenburg, D. J. Adams and A. I. Cooper, *Angewandte Chemie, International Edition*, **2016**, 55, 1824-1828

The authors regret that incorrect data was presented in Figure 2, Figure 3, and Table 1 of this publication.^[1] The corrected Figures and table are shown below. The hydrogen evolution rates were incorrectly calculated, but by a common scaling factor. Hence, the trends observed between materials and the overall conclusions made in the Communication remain valid. The correct H₂ evolution rate for the most active polymer, P7, under visible light (>420 nm) should be 37.3 μmol h⁻¹ (1492 μmol h⁻¹), not 92.0 μmol h⁻¹ as initially reported. The apparent quantum yields at 420 nm for P1K, P6, and P7 should be corrected to 0.4% (±0.1%), 2.2% (±0.2%), and 7.2% (±0.3%), respectively.

The most active polymer, P7, was studied independently by another research group,^[2] who reported an apparent quantum yield of 6.61%, close to our corrected value of 7.2%. We note that the precise value of the apparent quantum yield and hence the H₂ evolution rate will depend on the details of the experimental set up and the irradiation intensity.

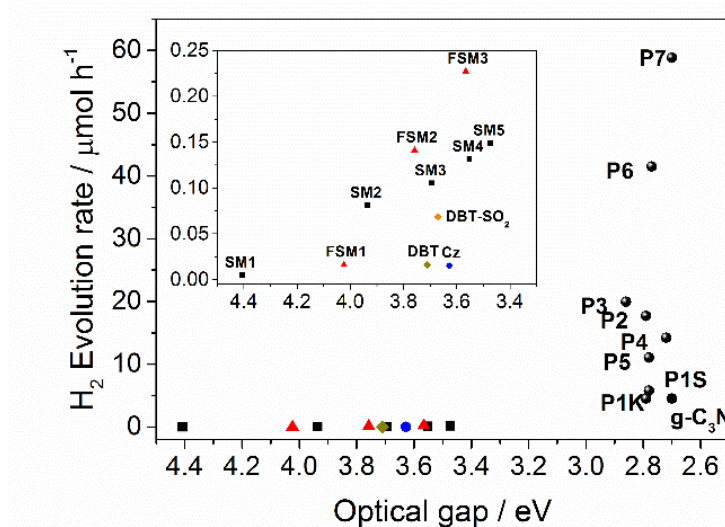


Figure 2. Photocatalytic hydrogen evolution rates. Each measurement was performed with 25 mg catalyst in water/MeOH/triethylamine mixture under broad-spectrum irradiation ($\lambda > 295$ nm; see Table 1 for visible light HERs).

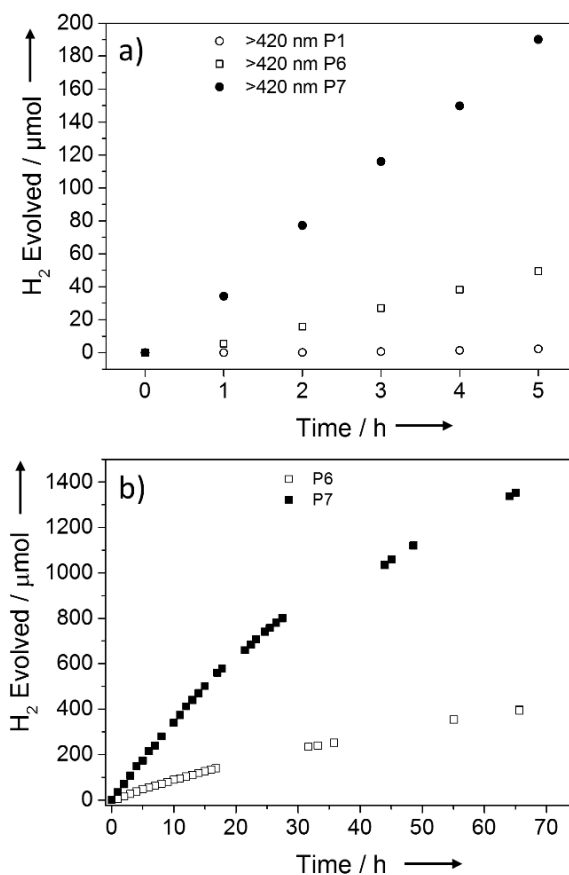


Figure 3. a) Time-course for photocatalytic H₂ production using visible light for P1K, P6, and P7 (25 mg catalyst in water/MeOH/triethylamine mixture $\lambda > 420$ nm); b) P6 and P7 (25 mg catalyst in water/MeOH/triethylamine mixture; $\lambda > 420$ nm), photolysis run for a total of 65 h.

Table 1. Photophysical properties and hydrogen evolution rates (HERs) for the polymer photocatalysts.

Polymer	Hydrogen evolution rate > 420 nm / $\mu\text{mol h}^{-1}$	Hydrogen evolution rate > 295 nm / $\mu\text{mol h}^{-1}$
P1K	0.8 (\pm 0.04)	4.2 (\pm 0.3)
P1S	1.6 (\pm 0.1)	5.8 (\pm 0.2)
P2	3.4 (\pm 0.1)	17.7 (\pm 0.1)
P3	>0.04 (\pm 0.02)	20.0 (\pm 0.2)
P4	3.2 (\pm 0.1)	14.2 (\pm 0.5)
P5	0.9 (\pm 0.2)	11.1 (\pm 0.2)
P6	10.8 (\pm 0.1)	41.5 (\pm 0.3)
P7	37.3 (\pm 0.8)	58.8 (\pm 1.9)

[¹] R. S. Sprick, B. Bonillo, R. Clowes, P. Guiglion, N. J. Brownbill, B. J. Slater, F. Blanc, M. A. Zwijnenburg, D. J. Adams and A. I. Cooper, *Angew. Chem. Int. Ed.*, **2016**, *55*, 1824-1828

[²] C. Yang, B. C. Ma, L. Zhang, S. Lin, S. Ghasimi, K. Landfester, K. A. I. Zhang, X. Wang, *Angew. Chem. Int. Ed.*, **2016**, *55*, 9202-9206.