



Short communication

## X-ray micro-tomography as a diagnostic tool for the electrode degradation in vanadium redox flow batteries



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### ABSTRACT

Micro-tomography (CT) can be successfully employed to characterize ex situ the structural changes occurring in graphite felt electrodes during vanadium redox flow battery (VRFB) operation. Coupled high resolution X-ray and electron microscopy in conjunction with XPS are used to elucidate the microstructural and chemical changes to the high voltage RFB carbon electrode. The results reveal the onset of corrosion of the carbon felt structure relatively early in the VRFB life-cycle, extended operation is expected to result in extensive microstructural evolution effects.

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### 1. Introduction

The vanadium redox flow battery (VRFB) is an electrochemical device suitable for large-scale energy storage due to its fast response time, long cycle life, and high energy efficiency [1,2]. In VRFBs, the electrodes provide the sites to facilitate the redox reactions between  $V^{3+}/V^{2+}$  and  $V^{4+}/V^{5+}$  ions [2,3].

Graphite felt is the most common electrode material employed in VRFBs due to its low cost, high stability and conductivity, and reasonable surface area [4,5]. To improve the activity of the felt, several surface treatments have been employed to alter the felt surface chemistry and, consequently, to improve the reaction kinetics at the electrodes. These include thermal activation [6], chemical etching/doping [1,7], nitrogen doping [8] and functionalization of the surface with hydroxyl groups [9]. Recent publications demonstrate that substantial performance gains can be achieved with improved design of these felt materials [10–13].

Tomograms of fresh graphite felt have already been obtained and used to extract detailed structural characteristics of electrodes (surface area, porosity, fiber diameter and volume) [14,15]. Based on the acquired structural information, 3D pore scale models have been built to investigate the effect of electrolyte flow rate, vanadium ion concentration and electrode morphology on VRFB performance. It has been found that overall cell voltage can be enhanced by increasing flow rate and by employing denser electrode structure [14,15].

However, even though this model-based design approach has helped optimize the fresh felt material, pore scale models for VRFBs would be much more accurate and realistic if the structural information of used (voltage-cycled) felts were incorporated in the model to capture the changes in felt structure and properties during operation.

Using a combination of micro-tomography (CT) measurements of voltage-cycled graphite felts (over ca. 30 charge/discharge cycles), scanning electron microscopy (SEM), and X-ray photoelectron spectroscopy (XPS) measurements of the same felt electrode, the first description of microstructural evolution processes occurring in VRFB felts during operation is provided in this paper.

### 2. Experimental

Graphite felts (Sigracell®; GFA6) were harvested from the positive electrode ( $V^{4+}/V^{5+}$ ) of a VRFB in a 'fresh' and 'used' state — 'fresh' felts were thermally activated and had undergone compression identical to those used in working VRFBs. 'Fresh' felts were in contact with the electrolyte solution, but had not been electrochemically cycled, whereas the 'used' felts underwent 33 charge/discharge cycles (Fig. 3(D)). Vanadium (IV) oxide sulfate (hydrate, 97%) was used as the starting electrolyte solution. In order to prepare  $V^{2+}$  and  $V^{5+}$  ions, anolyte and catholyte were filled with 1:2 ratio of V(IV) solution (1.5 M in 3 M  $H_2SO_4$ ) respectively and the required ions were obtained by charging of the VRB. The resulting solutions in anolyte and catholyte were 1.5 M  $V^{2+}/V^{3+}$  and 1.5 M  $VO^{2+}/VO_2^+$  respectively, both in 3 M  $H_2SO_4$  solution. Charge–discharge tests were conducted at a constant current density of 100 mA  $cm^{-2}$  at 30 °C. The cut-off voltages for the

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charge (upper limit) and discharge (lower limit) cycles were 1.65 and 0.65 V, respectively. The maximum cut-off voltage used during charging (1.65 V) was selected based on the maximum potential the battery could hold without excessive hydrogen evolution. The lower cut-off cell voltage during discharge (0.65 V) was selected to allow complete discharge of the battery.

Samples of 'fresh' and 'used' graphite felts were characterized in 3D using the micro-CT facility at the TOMCAT beamline at the Swiss Light Source. Using a parallel beam setup at 10.5 keV in absorption mode, 1500 projections were acquired for each tomographic scan during a 180° rotation with a radiation exposure time of 200 ms for each projection image, providing an effective pixel size of 0.65  $\mu\text{m}$ . Tomographic reconstruction of the acquired projection images was performed using the Gridrec algorithm [16] after application of flat and dark field corrections. Image analysis of the raw data was performed using the Avizo Fire software package (FEI): for each acquired tomographic volume, a cubic region-of-interest (0.48  $\text{mm}^3$ ) was extracted for subsequent analysis, a smoothing filter was then applied to the grayscale image sequences for image de-noising; followed by a watershed segmentation to separate the solid phase (graphite fibers) from the pore phase, this is a well-established work-flow that has previously been applied to graphite materials [17].

### 3. Results and discussion

Cross-sections of the micro-tomograms of 'fresh' and voltage-cycled graphite felts are shown in Figs. 1(a) and 2(a). 'Fresh' graphite felt fibers have uniform density, whereas an alteration of the structure of the used voltage-cycled graphite felts is observed, this micro structural evolution

is manifested as the appearance of agglomerated fiber bundles, and non-uniform density across individual fibers, consistent with electrochemical oxidation of the carbon structure.

The observed structural changes were verified by the 3D reconstructed images of graphite felt samples. 'Fresh' graphite felt (Fig. 1(B)) consists of individual fibers (fiber size  $\sim 13 \mu\text{m}$ ) forming 'bundles' throughout its structure, whereas the voltage-cycled felt (Fig. 2(B)) structure is dominated by agglomerated fibers.

Key geometric characteristics of the graphite felt samples (porosity, volume specific surface area and characteristic tortuosity) can be extracted from the 3D reconstructed images. Tortuosity values were calculated using an established methodology in StarCCM<sup>+</sup> software (see [18]). A small decrease in the calculated porosity and increase in tortuosity values [18] is observed in the case of voltage-cycled graphite felt as a consequence of the charge/discharge cycling for a short period of time ( $\sim 65 \text{ h}$ ). A small increase in the pore phase characteristic tortuosity of voltage-cycled electrode is observed due to the densification of the structure as a result of graphite felt oxidation. However, there is a significant reduction ( $\sim 37\%$ ) in the volume specific surface area of the voltage-cycled graphite felt sample after 60 h of continuous VRFB testing, indicating its structural alteration due to carbon oxidation and fiber agglomeration (Table 1). BET was employed to experimentally verify the observed changes in surface area by CT. The BET surface area of the 'fresh' sample was  $23 \text{ m}^2 \text{ g}^{-1}$  and it was reduced to  $14 \text{ m}^2 \text{ g}^{-1}$  for the voltage cycled sample, indicating a similar reduction in the surface area of the voltage-cycled graphite felt sample. This significant change in the microstructure of the felt will influence the microscopic transport and reaction phenomena occurring in the electrode, and thereby affect device-level performance.

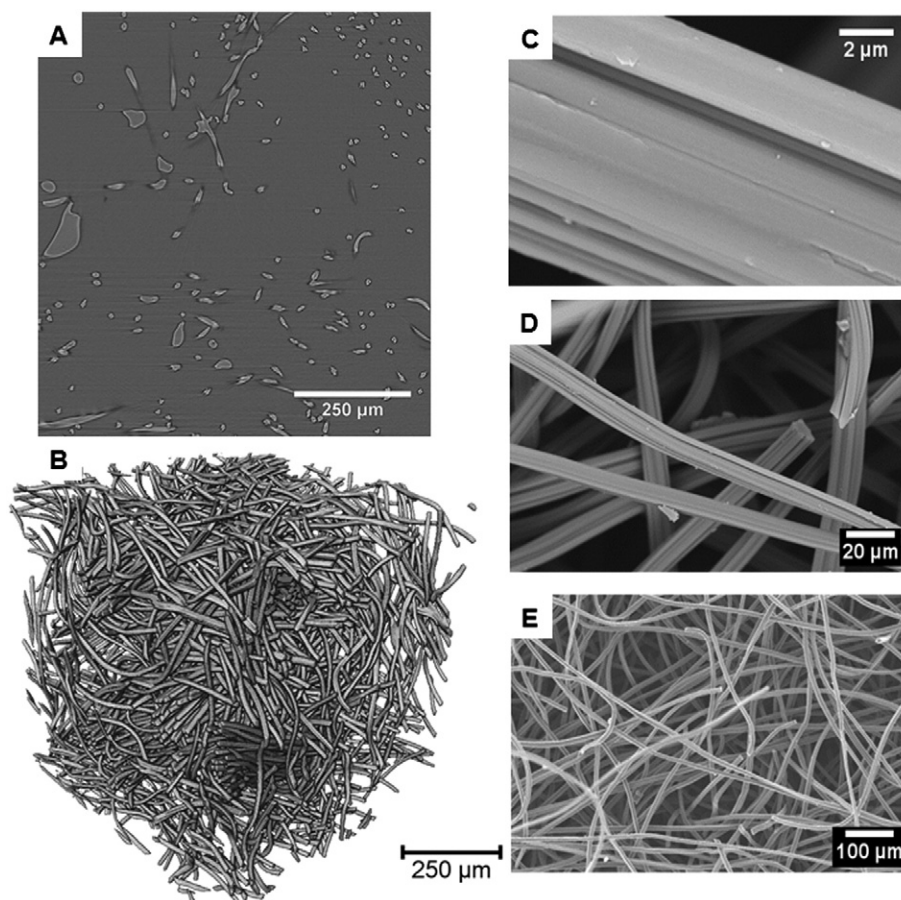
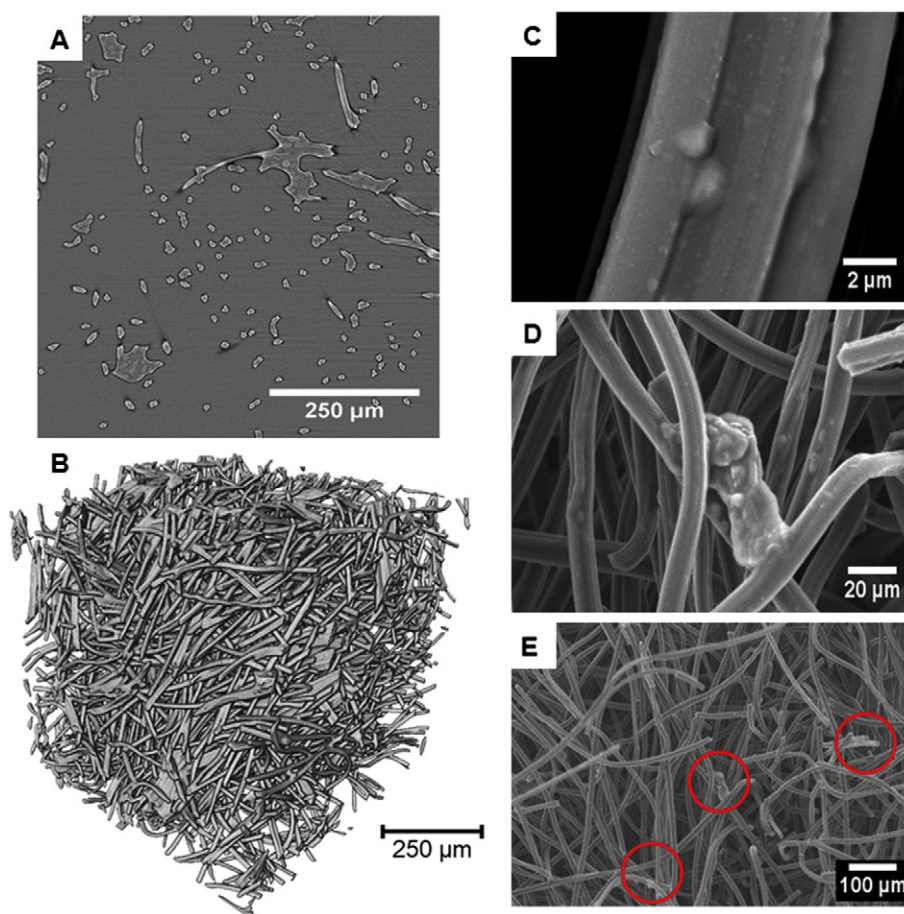


Fig. 1. A) Cross-section of the 'fresh' graphite felt electrode where the gray and dark gray regions represent the fibers and open pore structure respectively; B) 3D reconstructed image of the 'fresh' graphite felt electrode; c–e) SEM images of the 'fresh' graphite felt electrode.



**Fig. 2.** A) Cross-section of the voltage-cycled graphite felt electrode where gray and dark gray regions represent the fibers and open pore structure respectively; B) 3D reconstructed image of the voltage-cycled graphite felt electrode; C–E) SEM images of the voltage-cycled graphite felt electrode.

These structural changes observed using micro-CT have been verified using SEM, while XPS was used to investigate the changes on the surface of voltage-cycled electrodes.

SEM images of ‘fresh’ graphite felt (Fig. 1(C)–(E)) show that the individual fibers (~10 μm diameter) comprise bundles of several smaller fibers (~3–4 μm diameter each). The appearance of a few small cavities on the surface (Fig. 1(C)) was induced by thermal oxidation of the ‘fresh’ sample at 400 °C for 30 h to increase the surface area of the graphite felt [19,20] prior to VRFB charge/discharge cycling. Moreover, the bundle structure of the graphite felt results in edge planes and high defect concentration [21], and hence it is expected that the graphite felt fibers will react with oxygen and rapidly form highly oxidized carbon-oxygen groups [21,22]. SEM images of voltage-cycled graphite felts confirm this expectation by revealing layered oxidized graphite fibers beginning to agglomerate.

XPS studies have also been conducted to determine the surface chemical state of graphite felt before and after voltage cycling. The oxygen-to-carbon molar ratio increases from 0.09 to 0.39 in the case of the voltage cycled graphite felt; this pronounced ~300% increase in

surface oxygen content confirms that electrochemical oxidation of carbon occurs on the surface.

The C 1 s spectra for both ‘fresh’ and voltage-cycled graphite felts are asymmetric, with a tailing at the high energy binding side, which is typical for highly graphitized carbon fibers [21,23]. The C 1 s spectra of voltage-cycled graphite felt confirms the formation of COOH (~288.6 eV), a carbon oxidation intermediate [24,25].

Additionally, the O 1 s spectra of voltage-cycled graphite felts demonstrate that there is a considerable increase in the carboxylic surface groups (C=O and/or C–O–C), accompanied by a decrease in C–OH surface concentration; ~76% C=O and 24% C–OH present on voltage-cycled graphite felt compared to ~60% C–OH and 20% C=O present on ‘fresh’ graphite felt. These results suggest that the graphite felt has a strong tendency to interact with oxygen and form highly oxidized C=O species. Moreover, the observed C–OH reduction has an effect on the overall electrochemical reaction activity of the electrode as C–OH species enhance the absorption of vanadium ions on the surface of the electrode and catalyze the redox reaction between V<sup>4+</sup> and V<sup>5+</sup> ions [6]. The charge and discharge processes at the positive electrode involve oxygen transfer from the C–O–H bond to VO<sup>2+</sup> forming a surface VO<sub>2</sub><sup>+</sup> that diffuses back to the solution (Eq. (1)) [6,26].

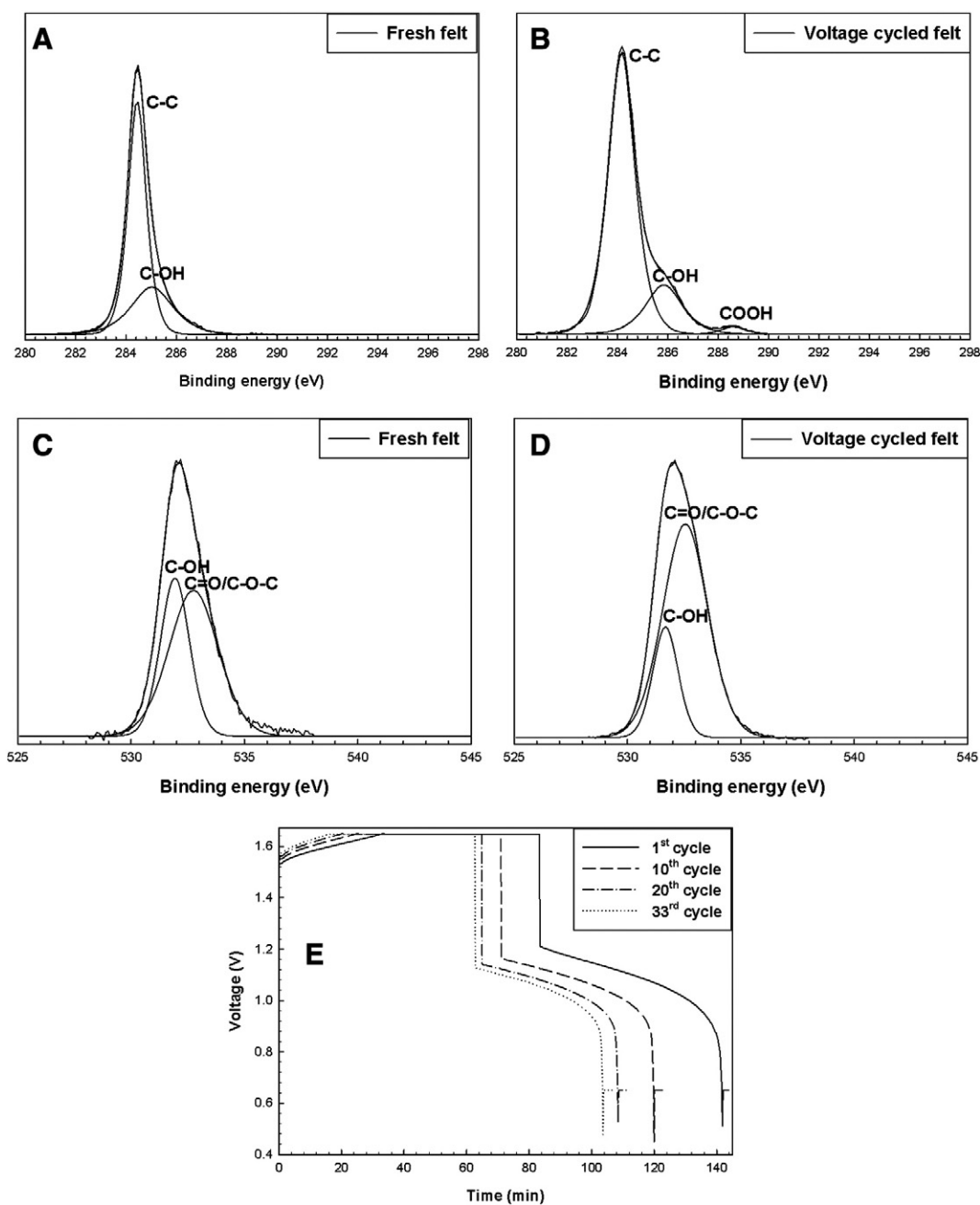


Thus, XPS results confirm that voltage-cycled graphite felt undergoes electrochemical oxidation during operation, resulting in the formation of carbon oxidation products (COOH) on the electrode

**Table 1**  
Geometrical properties of 3D reconstructed graphite felt electrodes.

	‘Fresh’ felt	Voltage-cycled felt
Porosity	0.941	0.924
Volume-specific surface area (μm <sup>-1</sup> )	0.495	0.328
Characteristic tortuosity [18]	1.014	1.026





**Fig. 3.** A) C 1 s spectra of 'fresh' graphite felt electrode; B) C 1 s spectra of the voltage-cycled graphite felt electrode; C) O 1 s spectra of 'fresh' graphite felt electrode; D) O 1 s spectra of the voltage-cycled graphite felt electrode; E) VRFB charge/discharge cycles (two laminated Nafion® 211 membranes as separator, current density:  $100 \text{ mA cm}^{-2}$ ,  $30^\circ \text{C}$ ).

surface and the reduction of surface COH species. The presence of C=O and COOH catalyzes the formation of  $\text{CO}_2$  accelerating the oxidation rate of the electrode [25,27].

Fig. 3(E) shows the charge–discharge curves obtained on a VRFB assembled with two laminated Nafion® 211 membranes (similar thickness to Nafion® 112) as separator. The data were obtained within a voltage window of 0.65–1.65 V at  $100 \text{ mA cm}^{-2}$  current density: preliminary charging tests at higher cut-off voltages resulted in the formation of abundant hydrogen bubbles at the negative electrode at cell voltages above 1.7 V. The 1st cycle has a lower charge voltage plateau and higher discharge plateau than the subsequent cycles, indicating that the voltage efficiency of the VRFB is higher before electrochemical oxidation of the electrode occurs. Voltage efficiency is gradually reduced as the number of charge/discharge cycles increases. An overall 30% capacity reduction

is observed, while the retention time of the last discharge (33rd cycle) is also decreased due to electrolyte and solvent transfer from the positive to the negative half-cell [28–31].

#### 4. Conclusions

Micro-CT is an exceptional diagnostic tool for monitoring the structural changes occurring in graphite felt electrodes during VRFB operation. Reconstructed 3D images of the VRFB electrodes reveal fiber agglomeration and carbon electrochemical oxidation during continuous battery operation. Key geometric characteristics of the graphite felt samples can also be extracted from these 3D images allowing the calculation of porosity, characteristic tortuosity and volume specific surface area (and the changes therein) for each sample. In this study, a significant decrease ( $\sim 37\%$ ) in the volume specific surface area of

voltage-cycled graphite felts is observed after only 65 h of continuous VRFB testing, indicating its structural alteration due to carbon oxidation/fiber agglomeration. SEM and XPS are employed to verify the structural and surface changes observed by micro-CT. SEM shows fibers with bundle structure starting to agglomerate after VRFB cycling and XPS confirms the electrochemical oxidation of graphite fibers, demonstrating the formation of an intermediate carbon oxidation product (COOH) on the electrode surface.

### Conflict of interest

The authors state that there is no conflict of interest.

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