

Fluorinated Metal–Organic Coatings with Selective Wettability

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ABSTRACT: Surface chemistry is a major factor that determines the wettability of materials yet devising broadly applicable coating strategies that afford tunable and selective surface properties required for next-generation materials remains a challenge. Herein, we report fluorinated metal–organic coatings that concurrently display water-wetting and oil-repelling characteristics, a wetting phenomenon different from responsive wetting induced by external stimuli. We demonstrate this selective wettability with a library of metal–organic coatings using catechol-based coordination and silanization (both fluorinated and fluorine-free), enabling sensing through interfacial reconfigurations in both gaseous and liquid environments, and establish a correlation between the coating wettability and polarity of the liquids. This selective wetting performance is substrate-independent, spontaneous, durable, and reversible, and occurs over a range of polar and non-polar liquids (60 studied). These results provide insight into advanced liquid–solid interactions and a pathway toward tuning interfacial affinities and realizing robust, selective superwettability according to the surrounding conditions.

INTRODUCTION

Coating materials that display specific interfacial affinities have attracted extensive interest, especially in the fields of chemical production, adhesives, selective membranes, bio–nano interactions, and repellent materials.^{1–6} The wettability of non-textured coatings is governed by surface chemistry,^{6–8} and chemically homogeneous surfaces typically have defined wettability—either repellent or wetting for a given liquid. Lower surface tension liquids (e.g., oil) will wet better than higher surface tension liquids (e.g., water) for nearly all surfaces.^{8–10} This phenomenon largely explains the challenge in creating surfaces with both water affinity and oil repellence properties.¹⁰ Nevertheless, achieving simultaneous controlled surface affinity^{11–15} and repellence to various liquids is important for engineering advanced materials¹⁶ and holds promise and broad interest for a wide range of wetting-related applications such as self-cleaning, chemical sensing, and

membrane separations.^{17–32} Externally induced wetting, through e.g., shear and strain^{33–35} or by combining surface active additives (e.g., chromophores),^{36–43} can initiate responsive wettability.^{44–48} However, externally induced wetting strategies do not lead to selective wetting without the application of specified external stimuli/energy or super-repellence to conventional wetting oils, in air or under water (e.g., droplet bouncing). Therefore, to date, a mechanistic understanding of selective wetting is largely unexplored. Furthermore, the development of broadly applicable coating strategies to achieve selective wettability remains a challenge.

Herein, selective and responsive wettability due to the surrounding fluid^{49–51} are achieved through dynamic and reversible surface rearrangement of the nanocomposite coatings that integrate active motifs with diverse chemical affinities, i.e., covalent binding of a highly non-wetting material atop highly wetting metal–organic thin films

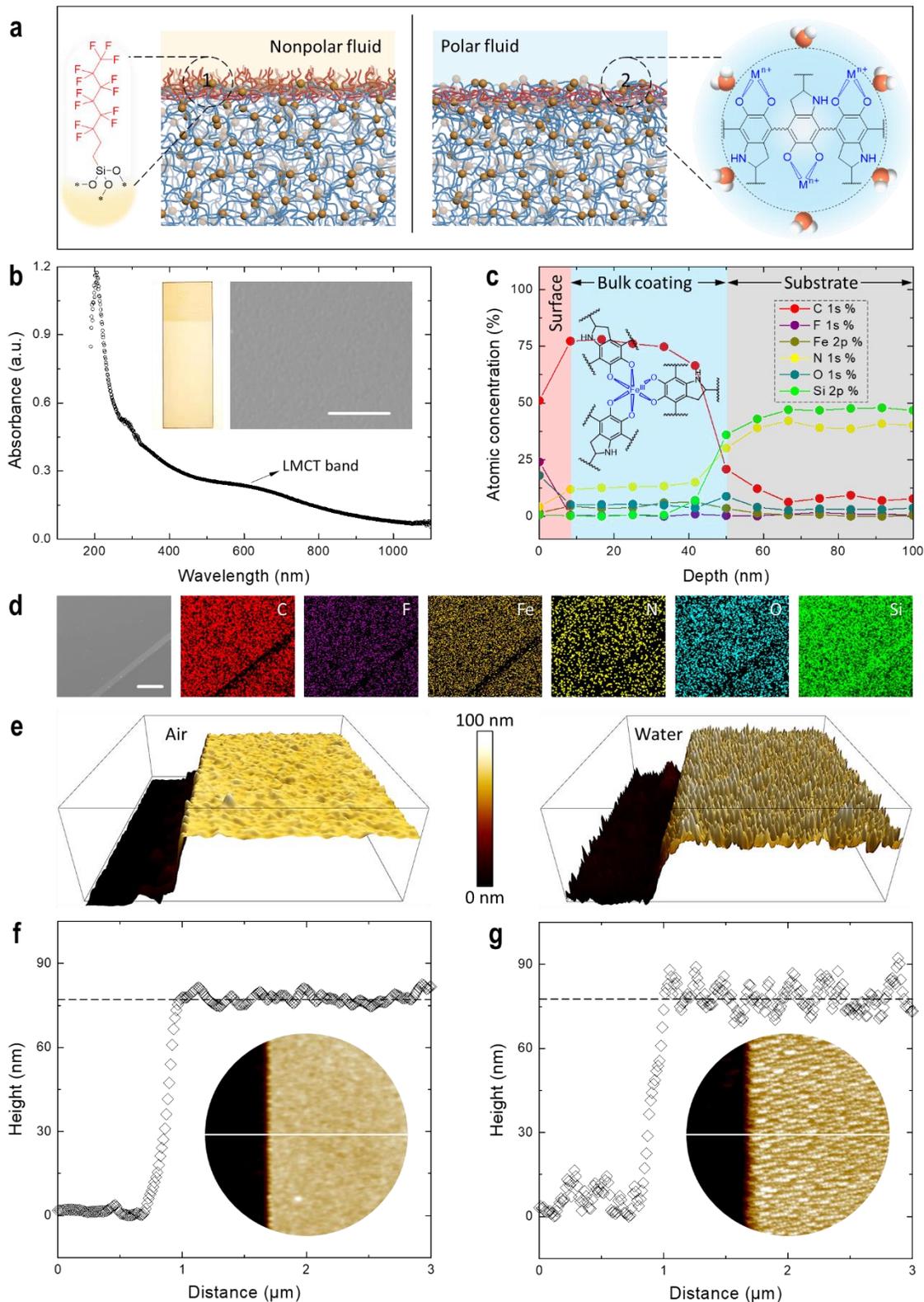


Figure 1. Design of repellent but polar heterogeneous nanocoatings. (a) Fluorinated metal-organic coatings (fMOCs) selectively respond to surrounding fluids according to their polarity and solvation ability, with a schematic of hydration (i.e., H_2O molecules) of the composite (M^{n+} represents metal ions) illustrated in the right panel. (b) UV-Vis spectra of fMOCs (fluorosilanzed Fe^{III} -polydopamine coating). The insets show a photograph of the corresponding coated glass ($2.6 \times 7.6 \text{ cm}^2$) and associated scanning electron microscopy (SEM) image. Scale bar is 2 μm . (c) Chemical composition of fMOCs on a silicon nitride substrate, as determined by XPS. The inset illustrates a possible building block of the bulk coating. (d) EDS mapping of a scratched fMOC showing the elements present on the surface. Scale bar is 50 μm . (e) 3D AFM images of a fMOC ($3 \times 3 \mu\text{m}^2$) in air and water. (f, g) Line profiles and 2D AFM images (insets) of a fMOC in air (f) and water (g). The black regions in (e-g) represent the scratched areas.

(Figure 1a). The nanocomposite coatings were engineered via the bio-inspired polymerization of metal–dopamine clusters⁵² and subsequent surface modification (silanization) (Figures S1 and S2). We used two types of trialkyl silanes (i.e., 1*H*,1*H*,2*H*,2*H*-perfluorooctyltriethoxysilane and its fluorine-free counterpart *n*-octyltriethoxysilane) as the precursors for surface silanization, as longer chain perfluorinated compounds (e.g., perfluorooctanoic acid or PFOA, perfluorooctanesulfonic acid or PFOS) are known to be toxic.^{53–58} The metal–organic coatings (MOCs) examined have been selected because the dopamine component provides universal adherence and affinity to most common liquids (e.g., polar liquids), the metal–chelation networks are chemically stable, active and modular, and tunable repellence to various liquids is achieved through the integrated low energy silane component and the high energy MOC component which can respond (i.e., flip-flop^{45,47} of the silane molecules) to the surrounding liquids based on their polarity (i.e., *f*MOCs; see *Materials and Methods* in the Supporting Information). Therefore, the resulting surface chemistry is heterogeneous, where both low surface energy silanes and high surface energy MOCs coexist in close proximity. In addition, the MOCs are intrinsically dynamic (or adaptive) owing to the nature of the coordination chemistry (between metal ions and the catechol moieties) exploited in the MOCs (i.e., reorganization of these components at the surface may occur).^{59–61}

RESULTS AND DISCUSSION

The heterogeneous nature of *f*MOC chemistry was studied using conformal coatings assembled by Fe^{III}–dopamine clusters, evident from the ligand-to-metal charge transfer (LMCT) band at ~600 nm (Figure 1b). These coatings were transparent on diverse substrates (e.g., glass, silicon wafer, polytetrafluoroethylene, and polypropylene) and allowed for chemical modifications including the chemical deposition and condensation of silane molecules (Figure S3). The X-ray photoelectron spectroscopy (XPS) and depth profiling of a ~50 nm *f*MOC reveals the distribution of the active components (i.e., fluorine and MOCs) from the surface through to the silicon nitride substrate. The fluorine motifs are only detectable within the superficial layer (i.e., during the first etch, ~8 nm), while the coordination chemistry is identical throughout the MOC layer (i.e., ~3 organic ligands per metal center) (Figure 1c). Although all elements are present on the surface, as revealed by energy-dispersive X-ray spectroscopy (EDS) elemental mapping (Figure 1d), the ratio between the repellent and attractive motifs (i.e., F/O ratio) was, on average, two (Figure 1c). The abundance of O species on the *f*MOC surface (due to e.g., adsorption, oxidation, solvation and/or preservation from MOC; Figures S4 and S5) coexisting at a comparable coverage (~30 at.%) with fluorine species that are present as a thin surface layer (<10 nm; Figure 1c) not only suggest the heterogeneous chemistry nature of *f*MOC, but also when combined with the amorphous nature of the coating could allow dynamic interfacial rearrangement and polar interactions with the MOCs.

In liquid, surface reconfiguration of *f*MOC, confirmed by atomic force microscopy (AFM), was more discernible when compared with that in air (Figure 1e). The root-

mean-square (RMS) roughness of the *f*MOCs following immersion in water (~4 nm) was more than twice as large as that of the dry *f*MOCs (~1.5 nm). This change is unlikely due to solvent swelling of the *f*MOCs (as the mean thickness of the film remained unchanged) but more likely due to the reconfiguration of the film occurring at the chemically heterogeneous surface (Figure 1f and 1g). This reconfiguration not only provides a blueprint for engineering selective wetting but also expands the toolbox of reconfigurable materials.

Surface Interactions and Wettability. Lateral force microscopy (LFM) investigations offered insight into the structural basis of reconfigurations observed in *f*MOCs (Figure 2a), provided their surface roughness was largely <1 nm in air (i.e., microscopically smooth, thereby enabling the study of the intrinsic interfacial frictions with given probes). As the LFM probes were rendered hydrophilic upon treatment with UV–O₃ for 10 min, higher signals in LFM images represent stronger hydrophilic interactions.⁶² Hydrophilic (blue) domains with irregular nano-shapes that coexisted with the repellent (red) regions for *f*MOCs (Figure 2a) were largely preserved from the bulk MOCs (Figures S6 and S7), facilitating the selective and dynamic response of *f*MOC surfaces to water. Notably, conventional repellent surfaces (e.g., fluorosilanized silicon wafer, *f*Silicon; Figure 2b) displayed negligible differences, confirming their homogeneous repellence nature (similar to Teflon), which is also evident from the color clustering analysis of the LFM images (Figures S7 and S8). The LFM experiments show that the *f*MOC is intrinsically polar despite being modified by non-polar silane molecules. Surface adhesion force measurements were also performed to demonstrate the difference between a conventional fluorinated surface (*f*Silicon) and *f*MOCs (Figure 2c). The *f*MOCs featured a snap-off force with the hydrophilic probe when retracting the probe from the surface in contrast to the repulsive force profile observed for *f*Silicon. It is noted that this non-conventional surface behavior was also observed for an alkanesilane-modified MOC surface (*a*MOC) (Figure S9), highlighting the generalization of the proposed design principle and its potential for engineering fluorine-free (thereby environmentally friendly) surfaces with selective wettability, although fluorocarbons have been reported as suitable platforms/interface materials for studying biosystems.⁶³ These LFM results and atomic adhesion force measurements not only highlight the static interfaces of *f*Silicon (i.e., non-responsive) compared with the adherent and responsive interfaces of *f*MOC (Figures S6–S9), but also emphasize the synergistic interplay and roles of the different species for the selective wettability of *f*MOCs.

The selective wetting of these *f*MOCs was also examined through the spreading dynamics of droplets (i.e., water, *n*-hexadecane) (Figure 2d). Oil droplets displayed virtually no spreading (contact angle ~75°), as predicted assuming homogenous surface chemistry (Movies S1 and S2). In contrast, the water droplet spread quickly (<1 s) until equilibrium (~23°), which is contrary (or counterintuitive) to what is expected for a typical repellent surface (*f*Silicon) (Movies S3 and S4). This observation is likely due to the interfacial reconfiguration of the *f*MOCs or the possible conformation/distribution changes of the surface additives.⁶⁴ Surface energy analysis provides a mechanistic

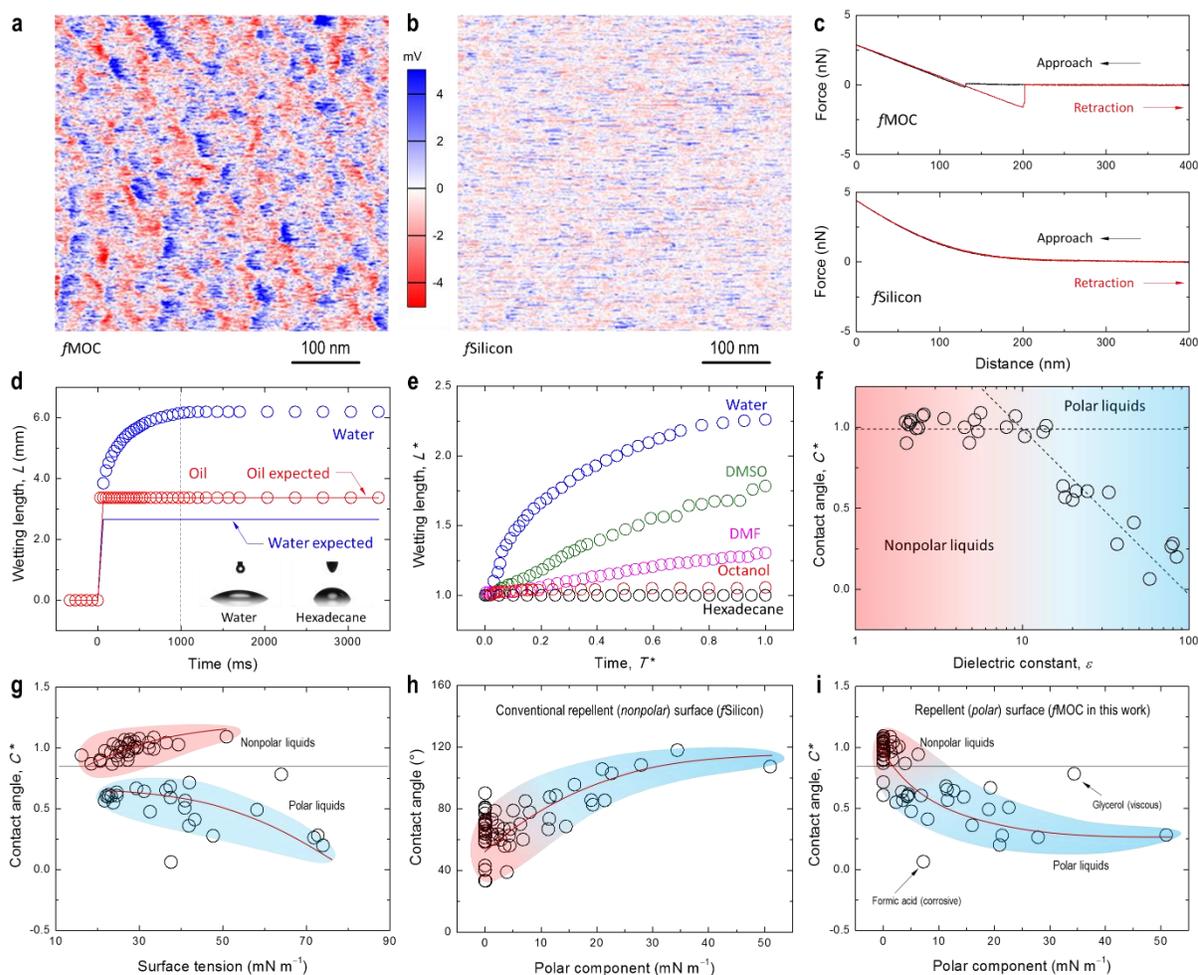


Figure 2. Selective water affinity and oil repellence. (a) LFM image of $fMOC$ ($0.5 \times 0.5 \mu m^2$), showing microscopically hydrophobic (red) and hydrophilic (blue) domains. (b) LFM image of $fSilicon$ control ($0.5 \times 0.5 \mu m^2$). (c) Representative atomic adhesion force curves of $fMOC$ and $fSilicon$ determined by colloidal-probe AFM. (d) Dynamics of the wetting length (diameter of the triple-phase contact line) of $\sim 5 \mu L$ liquids. Data of $fSilicon$ are presented by the (expected) lines. The insets are the silhouettes of water (left) and oil (right) droplets. (e) Wetting length (normalized with the droplet initial diameter) against the wetting time (normalized with the maximal wetting time). (f) Contact angle (normalized by conventional repellent control $fSilicon$) as a function of dielectric constant of the liquids. See Table S1 for liquids information. (g) Normalized contact angle of $fMOC$ as a function of surface tension of the liquids studied. (h, i) Contact angle of $fSilicon$ (h) and normalized contact angle of $fMOC$ (i) as a function of polar component of the liquids studied. The red lines and shaded areas (g-i) are visual guidelines. The black lines (g and i) serve as an indicative boundary between polar and non-polar liquids, in contrast to their non-selective control (h). Indicated outliers are due to their corrosive nature or viscous dissipation.

understanding of the selective wettability, showing the interplay of surface wetting and non-wetting forces. The $fMOCs$ have substantially different surface energy components (i.e., heterogeneity), namely a surface repellent non-polar component (fluorine) and a hydrophilic polar component (MOC), in contrast to less polar (lower energy) materials such as Teflon, polypropylene, and polydimethylsiloxane. Another characteristic of the $fMOCs$ is the presence of repellent motifs (i.e., fluorosilane) on top of the hydrophilic and dynamic MOC, and a reasonably low degree (e.g., F/O atomic ratio ~ 2) of non-polar surface chemistry affords selective and responsive wettability determined by the interplay of interfacial energies.

Further experiments indicated that the selective wetting also occurred for a diverse range of liquids (including *n*-hexadecane, *n*-octanol, dimethylformamide, dimethyl sulfoxide, and water), with water and oil, respectively, resulting in the most and least wetting properties (Figure

2e) regardless of the nature of the substrate. The selective wetting phenomenon also applied to a series of $fMOCs$ containing different metal ions⁶⁵ with nanoscale thickness (20–100 nm), tunable roughness and wettability, as well as transient timescale of reconfiguration (e.g., time of spreading is tunable and mostly ~ 0.2 s). The selective and spontaneous wetting of polar liquids also provides insights into the heterogeneous surface regime of $fMOCs$ (i.e., high polar interactions) and its potential in liquid separation afforded by its non-wetting property to non-polar liquids (which are typically preferentially wettable owing to their low surface tensions). In addition, the different binding constants⁶⁶ between a given organic ligand (catechol) and different metal ions are expected to allow tuning of the stability of heterogeneous $fMOCs$.

To further examine the surface wettability, 32 liquids with different physiochemical properties (i.e., ϵ values ranging from 2 to 100) were examined (Figure 2f, Table S1). Liquid (as-placed or static) contact angles^{67,68} on

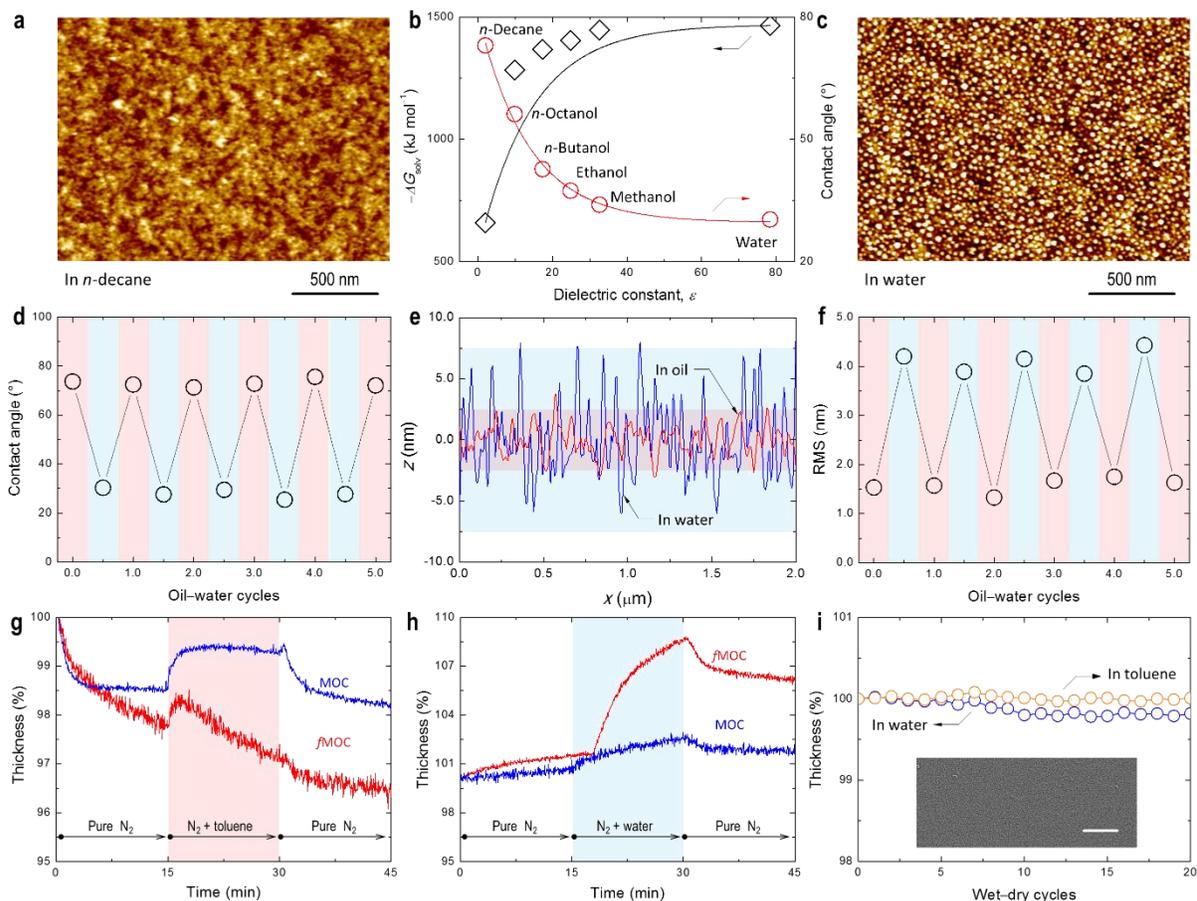


Figure 3. Reversible topology and wettability. (a) AFM image of *f*MOC in oil (*n*-decane). Scale bar is 500 nm. (b) Correlation between calculated solvation energy (diamond) of Fe^{III}-dopamine clusters and the experimental *f*MOC surface wettability (circle). (c) AFM image of the coating (a) in water. Scale bar is 500 nm. (d) Contact angles of oil (*n*-hexadecane) and water as a function of cycle number. The cyclic test involved successive contact measurements using a new droplet for each new cycle after removing the old droplet by blowing nitrogen. (e) Representative AFM line profiles of the *f*MOC in *n*-decane and water. (f) RMS roughness of the *f*MOC as a function of cycle number (in *n*-decane and water). (g, h) Thickness variation of MOC and *f*MOC thin films exposed to a (g) toluene atmosphere and (h) water atmosphere. (i) Durability of *f*MOC over 20 wet-dry cycles (either bulk oil or water was used to wet the coating via immersion for 5 min, followed by drying in nitrogen). The inset shows a helium ion microscopy image of the undamaged *f*MOC film even after 20 wet-dry cycles. Scale bar is 1 μ m.

*f*MOCs were normalized by their controls, excluding interfacial liquid confinement.⁶⁹ The surface was passive to all of the non-polar liquids studied but displayed selective affinity to the polar liquids, which can be also seen from the increased contact angle hysteresis of polar liquids, in contrast to the conventional repellent and non-polar surfaces (e.g., *f*Silicon) (Figure S10). To provide further insights into surface regimes, the wettability of 60 liquids (see *Probe Liquids in Materials and Methods* in Supporting Information) was plotted against their surface tension (consisting of polar (p) and non-polar (dispersive; d) components, i.e., γ_1^p and γ_1^d , respectively) (Figure 2g–2i). C^* is a measure of the extent that *f*MOC differs from the *f*Silicon control, i.e., conventional repellent (non-polar) surface. $C^* \sim 1$ indicates that *f*MOC has a comparable surface property with *f*Silicon control, whereas $C^* \sim 0$ indicates a complete reversed interfacial property, i.e., non-conventional repellent surface with preserved selective affinity. Contrary to the observation made for the conventional repellent surfaces (i.e., non-polar *f*Silicon, Figure 2h), the contact angles of the liquids residing on our polar *f*MOC clustered in two regions (Figure 2g) and, in general, coincided with the surface tension of their

predominant polar components (Figure 2i). For example, *f*Silicon follows the Young's equilibrium^{67,68} at the triple phase contact line of a droplet at the surface, i.e., $\cos \theta \sim (\gamma_1^p + \gamma_1^d)^{-1}$ for a given surface energy (Figure 2h). This principle applies for conventional surfaces but does not apply for the heterogeneous *f*MOCs that attract polar liquids strongly, as indicated by the two separate regions observed in Figure 2g. As for *f*MOCs, the relation between liquid contact angle and its surface tension components has to be considered separately as $\cos \theta \sim (\gamma_1^d)^{-1}$ for non-polar liquids and $\cos \theta \sim (\gamma_1^p)^{-1}$ for polar liquids. This decoupled relation together with C^* thereby allow prediction of surface wettability of heterogeneous surfaces and their degree of surface heterogeneity. This finding is consistent with other correlations determined from parameters used to evaluate liquid polarity including dipole moment,⁷⁰ permittivity,⁷¹ polarity index,⁷² and empirical parameters⁷³ (Figure S11), thereby demonstrating the heterogeneous nature of *f*MOC. It is noted that the selective wettability may also be partially ascribed to the zeta potential of *f*MOC (tunable within the range of -60 to 60 mV at pH 1–12; Figure S12) and/or the ionic strength^{74,75} but is not related to surfactant-driven

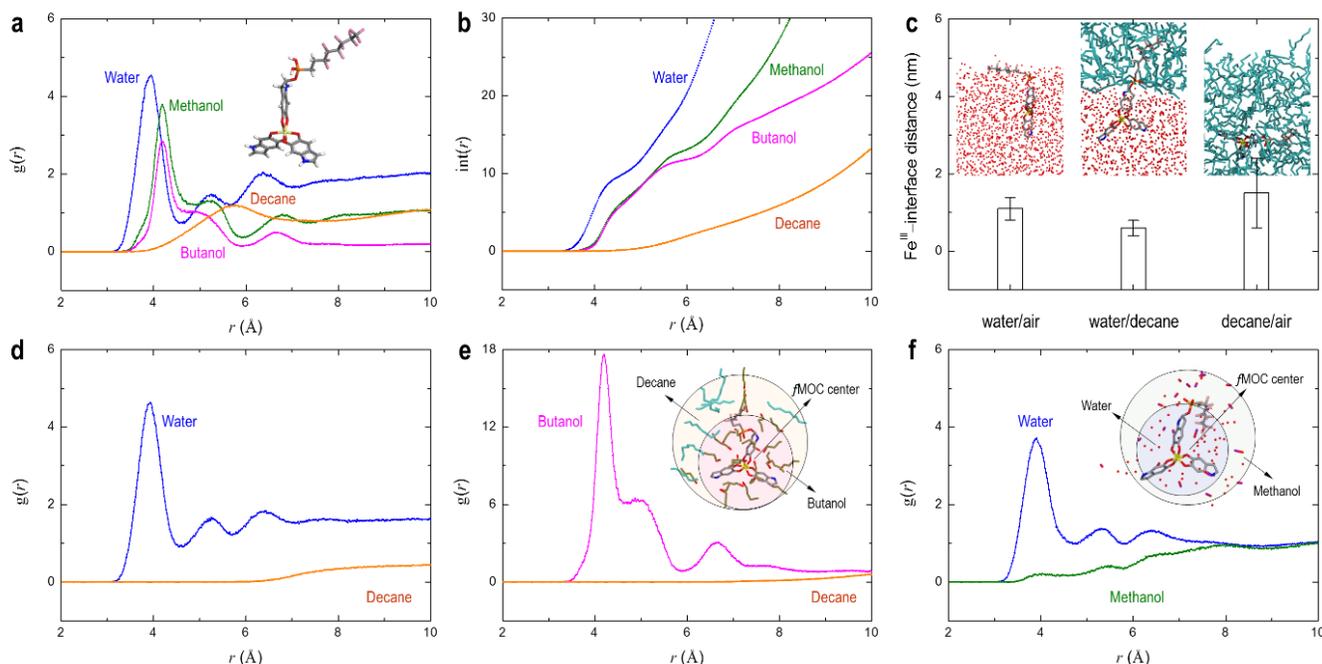


Figure 4. Molecular dynamics simulations of solvation shell and solvation selectivity of *f*MOC with liquids. (a, b) Radial distribution functions $g(r)$ (a) and integrals $\text{int}(r)$ (b) between Fe^{III} *f*MOC (inset) and different liquids (from polar to non-polar: water, methanol, butanol, and decane). (c) Distance from center of *f*MOC (e.g., Fe^{III}) to the interface and representative snapshots from MD simulations of the model *f*MOC subunit in water, 1:1 (v/v) water/*n*-decane (immiscible) mixture, and *n*-decane with a solvent-air interface. Five independent simulations for each system were generated and error bars are standard deviations. (d–f) Radial distribution functions $g(r)$ between Fe^{III} *f*MOC and oxygen (or carbon for decane) in three liquid mixtures (1:1 v/v): immiscible water/*n*-decane mixture (d), miscible organic (*n*-butanol/*n*-decane) mixture (e), and miscible aqueous (water/methanol) mixture (f). The insets show the corresponding solvation shell states.

transient wetting based on the unchanged liquid surface tensions before and after contacting the *f*MOCs (Table S2). This unaltered liquid property also indicates that *f*MOCs are chemically stable when contacting the probing liquids—a prerequisite for durable interface reversibility.

Interface Reversibility. In addition, alcohol homologs were chosen (i.e., *n*-octanol, *n*-butanol, ethanol, and methanol) to evaluate the potential solvation effects and morphology changes when exposed to different liquids (Figure 3a–3c). Previous studies^{49–51} have shown that some functional groups are more exposed to the surface than others as a function of the liquid above the solid. Our results also showed that the selective wetting is inversely correlated to the solvation energies of the Fe^{III} -dopamine clusters (i.e., the solvation free energy, ΔG_{solv} (see *Quantum Mechanics (QM) Calculations of Solvation Free Energy* in the Supporting Information), decreases with increase in contact angle) (Figure 3b). Importantly, the selective wetting was reversible, and no obvious change was observed after 100 measurement cycles (where one cycle refers to one iteration of exposure to oil and water, data of 5 cycles of measurements are shown in Figure 3d). This was apparent from the nanoscale surface configurations (e.g., height profile) of *f*MOCs in the corresponding liquids (Figure 3e), where control unmodified coatings showed negligible reconfiguration (Figure S13). In addition, we investigated the nanoscale surface reconfiguration and reversibility by measuring the surface roughness of the coatings when subjected to multiple oil–water immersion cycles (observing changes within a single cycle only was not feasible as these changes occurred within a short timescale of ~ 1 s). The surface roughness was largely

preserved over multiple cycles of AFM measurements (e.g., RMS in Figure 3f), with negligible changes over 100 repeated tests (i.e., intrinsically dynamic nature) (Figure S14). It is noted that the extent of surface reconfiguration is typically in line with the polarity of the surrounding liquids (Figure S15) and can be further altered by the incorporation of different metal ions. The selective wetting behavior of the *f*MOCs to surrounding environments also applies to gaseous conditions. Specifically, the thickness of the *f*MOCs was monitored in real time by ellipsometry during their exposure to dry nitrogen and solvent vapors (i.e., toluene, water) (Figure 3g and 3h). The variations in thickness revealed the specific interaction between the surface and the vapor. For example, the decrease in thickness observed in Figure 3g indicated the desorption of moisture from the *f*MOC surface when non-polar toluene vapor was introduced, whereas the increase in thickness in Figure 3h indicated the reconfiguration of *f*MOC due to the surrounding water vapor. The results highlight that the *f*MOCs are sensitive to water vapor (polar) but relatively inert to gas-phase toluene (non-polar). That is, the coating displays composition-dependent selective response to the gaseous environment. In contrast, the unmodified MOCs had negligible selectivity, which explains the importance of surface heterogeneity, which is present in *f*MOCs, for selective wetting. Notably, *f*MOCs are stable over 20 wet-dry cycles (Figure 3i) and after long-term (up to 10 days tested) immersion tests in water and oil (determined from the negligible changes observed in coating thickness, zeta potential, and refractive index), reversible morphology tests (Figure 3f, Figure S14), and liquid surface tension tests (Table S2).

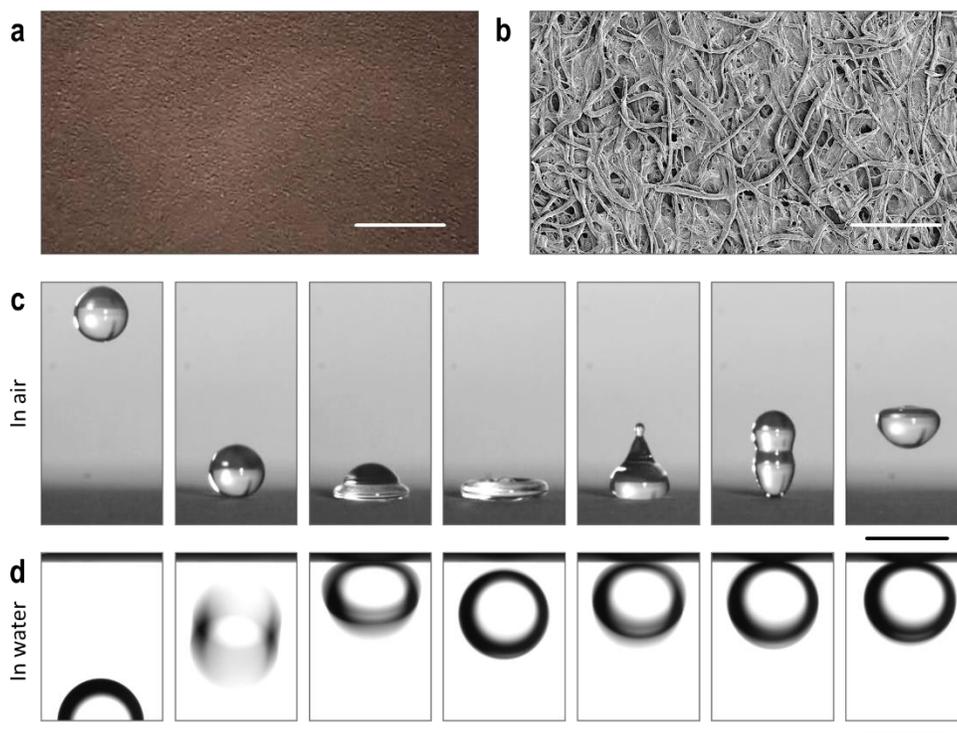


Figure 5. Robust super-repulsion to oil in either air or water. (a, b) Digital and SEM images of a coarse substrate (filter paper) after *f*MOCs treatment. Scale bars are 1 cm and 500 μ m, respectively. (c, d) Bouncing oil droplets (*n*-hexadecane) on the coated substrate either placed in air (c) or immersed in water (d). Scale bars are 5 mm. Time intervals are \sim 4 and \sim 10 ms, respectively.

Molecular Mechanistic Study. Molecular dynamics (MD) simulations were performed (see *Materials and Methods* in the Supporting Information) to obtain a further understanding of the interfacial interactions of the heterogeneous chemistry of *f*MOC and its selective wettability (Figure 4, Figure S16). By placing the subunit of *f*MOC coating (i.e., *f*MOC cluster) in different liquids with varying polarities, solvation can be represented by the radial distribution function peak maximum (i.e., solvation shell; Figure 4a) and the integrals (Figure 4b). It is noted that the subunit of *f*MOC simulated as an all-atom film simulation was not feasible but should provide a basic understanding of the heterogeneous *f*MOC structure. Molecules of polar liquids (e.g., water, alcohols) associate with *f*MOC within closer proximity when compared with non-polar liquid (e.g., *n*-decane) (Table S3), indicating the higher affinity of *f*MOC to polar liquids than non-polar liquids. In addition, the molecular mechanistic of the selective wettability of *f*MOC can be evaluated by MD simulations using liquids mixtures either immiscible or miscible (Figure 4c–4f). The *f*MOC–solvent interactions can be explained by their selective occurrence at the phase interface with the polar motif residing on the polar phase (Figure 4c) and the selective solvation in mixed (competitive) environments including immiscible water/decane mixture (Figure 4d), miscible and less polar butanol/decane system (Figure 4e), and miscible and polar water/methanol system (Figure 4f). These MD simulations collectively support the selective water-wetting–oil-hating behavior of the heterogeneous *f*MOCs.

Superwettability. The non-conventional wetting regime was further applied to and could be further enhanced by coarser substrates (Figure 5a and 5b). The selective wetting was examined in both air and liquid. For example, an oil droplet bounced off an *f*MOCs-coated filter paper

both in air (Figure 5c) and water (Figure 5d) without penetration, demonstrating super-repulsion to oil (Movies S5 and S6). The super-repulsion achieved (i.e., contact angle $>150^\circ$, roll-off angle $<10^\circ$, and droplet bouncing both in air and in liquid) can be ascribed to the reduced solid–liquid contact on coarser substrate (i.e., higher degree of structural hierarchy tiers)⁶ and also indicate the potential for tuning surface selective repulsion and affinity. Furthermore, these surfaces displayed selective super-affinity to a polar liquid (e.g., water) in both air and oil environments (Figure S17, Movies S7 and S8). Collectively, these findings demonstrate extreme instances of selective wetting and highlight how heterogeneous coatings can open new routes to studying fundamental questions in surface wettability.

CONCLUSION

A coating strategy for engineering selective surface wettability is developed by combing metal–organic chemistry and low-energy silanization. Both experimental and computational results reveal that the *f*MOCs are heterogeneous in nature and display an unexpected wetting mechanism: the coexistence of repulsion to low surface tension oils and wetting by polar and/or high surface tension liquids such as water. Unlike conventional repellent surfaces where the degree of surface repulsion (i.e., from high to low) follows the order of liquid surface tension (i.e., from high to low), the present heterogeneous *f*MOC surfaces display repulsion/wetting that follows the order of both liquid surface tension and their polarity. The heterogeneous *f*MOCs are capable of decoupling the interfacial polar and nonpolar interactions with quantitative correlations determined, providing a general pathway that allows for advanced and selective wettability engineering, e.g., the coexistence and decoupled oil

repellence and liquid wetting (for polar liquids) enables gravity-driven (highly energy efficient) liquid separations. Although further experiments are required for a complete understanding of the interplay between materials and fluids,^{18,67,76,77} the flexibility of designing tunable interfaces contributes to the knowledge of nature-inspired superwetting systems.²⁰ This highlights that the engineering of coatings with selective wettability is not only feasible but also promising for on-demand practical applications in a wide range of fields, e.g., selective liquid sieving, fog harvesting, oil spill clean-up and recycling, and environment-friendly cleaning of clothes and surfaces.

ASSOCIATED CONTENT

Supporting Information. Materials and methods (coating preparation, vapor sensing, surface adhesion, lateral force analysis, wettability and selectivity studies, quantum mechanics calculations, MD simulations, and characterization), Figures S1–S17 (surface chemistry, lateral force imaging, fluorine-free coatings, contact angle hysteresis, correlation between wettability and polarity, AFM roughness and responsiveness, MD simulations, and decoupled superhydrophilicity), Tables S1–S3 (properties of probing liquids and solvation data), and Movies S1–S8 (selective wetting/nonwetting and superwettability of coatings in air and in liquid). This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Notes

The authors declare no conflict of interest. All authors have given approval to the final version of the manuscript.

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ABBREVIATIONS

MOC, metal–organic coating; fMOC, fluorinated metal–organic coating; aMOC, alkanesilane-modified MOC; fSilicon, fluorosilanized silicon wafer; SEM, scanning electron microscopy; XPS, X-ray photoelectron spectroscopy; LMCT,

ligand-to-metal charge transfer; EDS, energy-dispersive X-ray spectroscopy; LFM, lateral force microscopy; AFM, atomic force microscopy; RMS, root-mean-square; MD, molecular dynamics.

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