



# Self-assembled carbon nanoribbons with the heteroatom doping used as ultrafast charging cathodes in zinc-ion hybrid supercapacitors

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**ABSTRACT** Zinc-ion hybrid supercapacitors (ZHSs) are highly desirable for large-scale energy storage applications owing to the merits of high safety, low cost and ultra-long cycle life. The poor rate performance of cathodes, however, severely hinders their application. Herein, aqueous ZHSs with superior performance were fabricated by employing a series of ultrathin carbon nanobelts modified with B, N, O (CPTHB-B<sub>x</sub>). The heteroatom doping can significantly modify the chemical behaviors of carbon frameworks, which could generate numerous active sites and accelerate the charge transport. The systematic investigation reveals that the B-N groups are active species for fast Zn-ion adsorption and desorption. As a result, the best-performed CPTHB-B2 exhibits an excellent electrochemical performance as cathodes in ZHSs, delivering a high specific capacitance of 415.3 F g<sup>-1</sup> at 0.5 A g<sup>-1</sup>, a record high capacitance retention of 81% when increasing the current densities from 0.5 to 100 A g<sup>-1</sup>, an outstanding energy density of 131.9 W h kg<sup>-1</sup> and an exceptionally high power density of 42.1 kW kg<sup>-1</sup>. Our work provides a new cathode design for ultrafast charging Zn-ion storage devices.

**Keywords:** carbon nanoribbons, self-assembled, heteroatom doping, ultrafast charging Zn-ion storage

## INTRODUCTION

Electrochemical energy storage devices, including batteries and supercapacitors, have attracted great attention due to the increasing demands of electronics, electric vehicles and intermittent renewable energies [1–3]. Lithium-ion batteries, featured with high energy density, exhibit safety issues and high fabrication cost owing to the activity of alkali metals and flammable organic electrolytes [4,5]. In contrast, zinc-ion electrochemical energy storage devices, including zinc-ion batteries (ZIBs) and hybrid supercapacitors (ZHSs), show a high theoretical capacity of 820 mA h g<sup>-1</sup> (Zn/Zn<sup>2+</sup>) and low redox potential (−0.762 V vs. standard hydrogen electrode) of Zn anodes as well as aqueous electrolytes offering high safety and low cost [6–8]. Currently, the research interest in cathode materials, Zn anodes and elec-

trolytes for ZIBs are increasing dramatically [9–12]. Nevertheless, the dissolution and parasitic reactions of transition metal oxide cathodes can lead to poor cycling stability and limit the wide application of ZIBs [13,14].

Currently, ZHSs, consisting of zinc anodes and porous carbon cathodes, can combine the advantages of secondary batteries and supercapacitors such as high working voltage and ultra-long cycle life, which are recognized as promising candidates for next-generation energy storage devices [8,15]. In 2018, Dong *et al.* [16] firstly reported aqueous ZHSs using an activated carbon cathode and a Zn metal anode, and provided new insights for Zn-based energy storage devices. The device delivered a specific capacity of 121 mA h g<sup>-1</sup>, a large energy density of 84 W h kg<sup>-1</sup> and power density of 14.9 kW kg<sup>-1</sup>. Recently, although various porous carbon cathodes are designed to enhance the performance of ZHSs, carbon cathodes still suffer from insufficient active sites and sluggish charge transfer processes, further leading to the low capacity and poor rate performance [17–19]. Consequently, the fabrication of carbon materials with abundant active sites and short charge transfer barriers are in urgent need to enhance Zn-ion storage capability of ZHSs cathodes.

It has been demonstrated that introducing heteroatoms into the carbon skeleton is an efficient strategy to solve these issues, which can provide more active sites and enhance the electrical conductivity [20–22]. Besides, the studies on developing suitable active structures and synergy between heteroatoms also need intensive exploration. Herein, we tackle these limitations by introducing heteroatoms (B, N and O) into ultrathin carbon nanobelts (CPTHB-B<sub>x</sub>), which effectively boosts the electrochemical performance for ZHSs. The heteroatoms can not only serve as active sites to enhance Zn<sup>2+</sup> adsorption, but also promote the charge transfer rate. Meanwhile, the doping amount of B atoms was controlled and the B-N group was identified as the major active sites as the Zn-ion hosts. As a consequence, the resulting CPTHB-B2 electrode exhibited an excellent specific capacitance of 415.3 F g<sup>-1</sup> at the current density of 0.5 A g<sup>-1</sup>, one of the highest rate capabilities (81% capacitance retention at 100 A g<sup>-1</sup>) and a superb stability (capacitance retention 98.5%) after 10,000 cycles at a current density of 10.0 A g<sup>-1</sup>. Further-

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more, an energy density of  $131.9 \text{ W h kg}^{-1}$ , together with a remarkable power density of  $41.2 \text{ kW kg}^{-1}$  were achieved. To demonstrate practical applications, a flexible solid-state pouch cell ZHS based on the CPTHB-B2 cathode was constructed, which delivered a good cycling performance under different bending states.

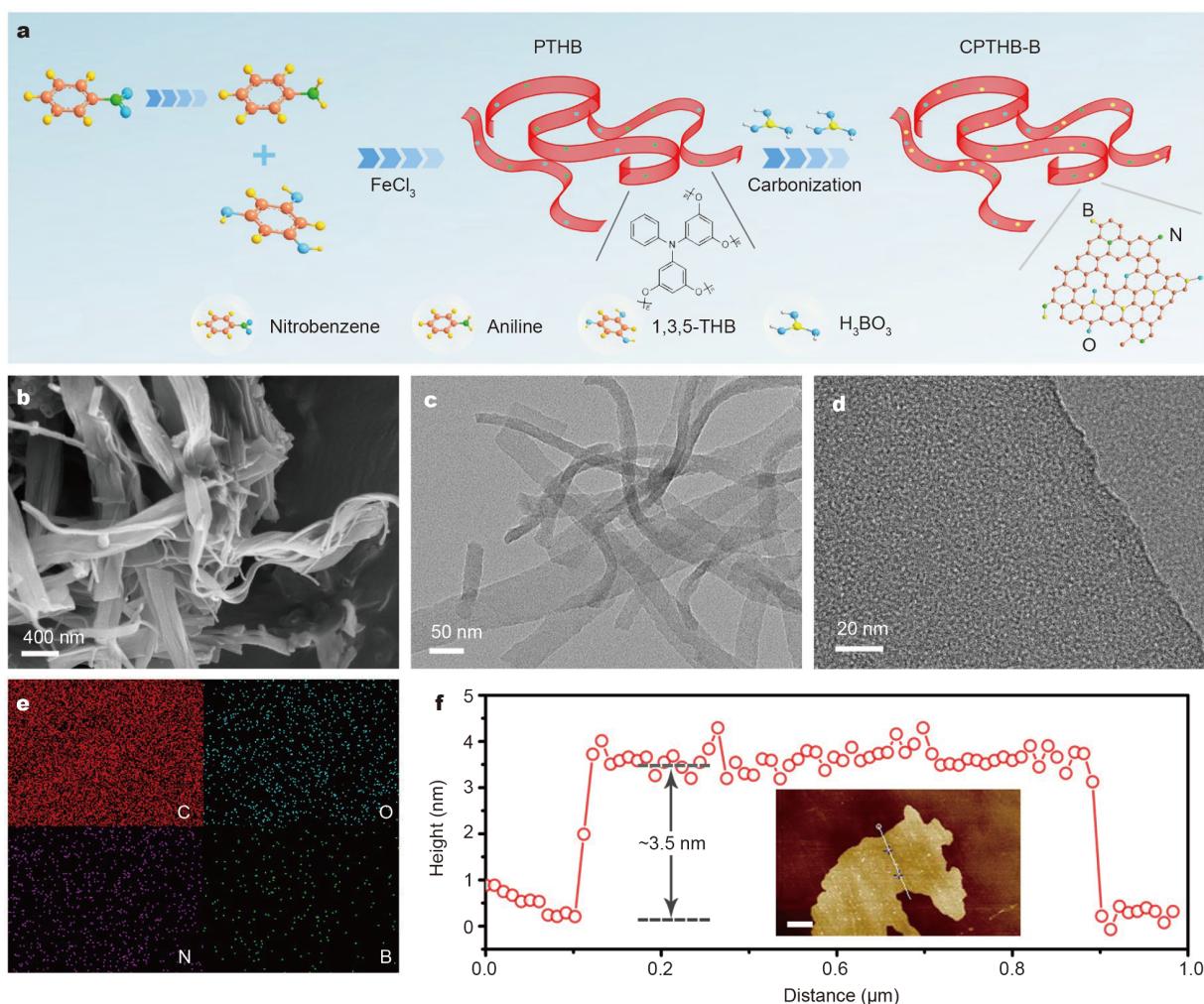
## RESULTS AND DISCUSSION

The preparation process of CPTHB-B<sub>x</sub> is illustrated in Fig. 1a. As for the formation of PTHB (Fig. S1), the nitrobenzene can not only serve as the solvent in the following polycondensation process, but also as the reactant after being reduced into aniline. The Fourier transform infrared (FTIR) spectrum of the PTHB (Fig. S2) shows an absorption band at  $1100 \text{ cm}^{-1}$  for stretching vibration of C–O–C, manifesting the successful synthesis of PTHB [23]. The as-prepared PTHB with N, O atoms doping and boric acid ( $\text{H}_3\text{BO}_3$ ) were mixed with specific mass ratios (PTHB:  $\text{H}_3\text{BO}_3 = 1:x$ ,  $x = 1, 2, 3, 4$ ). Then, the mixture was carbonized at  $800^\circ\text{C}$  under argon atmosphere, and the B atom was further introduced into carbon materials.

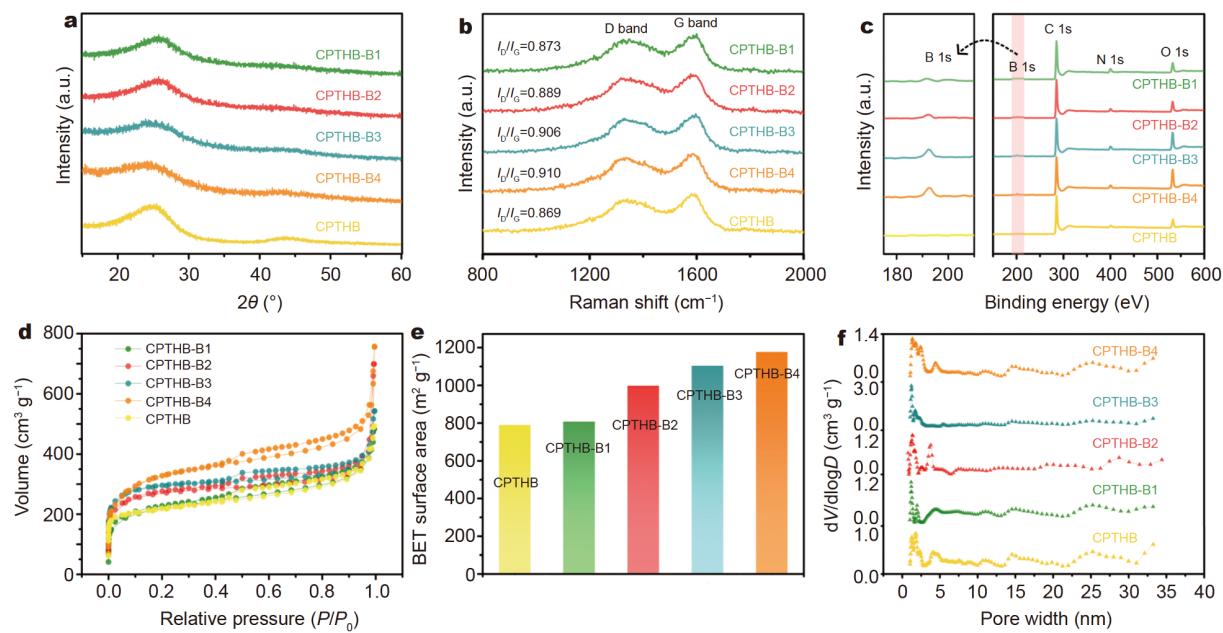
The morphologies and microstructures of CPTHB-B<sub>x</sub> and CPTHB samples were observed by scanning electron microscope (SEM), and the results are displayed in Fig. 1b and Figs S3, S4.

Fig. 1b shows that the CPTHB-B2 exhibits an ultrathin carbon nanobel structure, and the samples with other ratios show the same morphology (Fig. S3), which is further verified by transmission electron microscopy (TEM) in Fig. 1c. The structures of CPTHB-B<sub>x</sub> exhibit great similarities with CPTHB (Fig. S4), indicating the structure and morphology of nanobelts are well maintained after the introduction of B atoms. From the high-resolution TEM (HR-TEM) in Fig. 1d, the CPTHB-B2 is governed by disordered nanopores. In addition, the elements of C, O, N, and B are well-distributed in the nanobel structure, as depicted in the SEM mapping images (Fig. 1e and Fig. S5). The thickness of the CPTHB-B2 nanobelt was measured by atomic force microscope (AFM) and the randomly measured carbon nanobelts have a thickness of  $\sim 3.5 \text{ nm}$  (Fig. 1f). The ultrathin nanobel structures shorten the distance of ion mobility, thereby facilitating the charge transfer process.

To investigate the structures of the carbon nanobelts, we conducted X-ray diffraction (XRD) measurements of CPTHB and CPTHB-B<sub>x</sub>. Fig. 2a shows two distinct diffraction peaks at around  $25^\circ$  and  $43^\circ$ , which are indexed to the (002) and (100) lattice planes of graphite carbon, respectively [24]. The structural features were further measured by Raman spectroscopy. The Raman spectrum (Fig. 2b) shows a D band (representing dis-



**Figure 1** (a) Schematic illustration of the preparation of CPTHB-B. (b) SEM, (c) TEM, (d) HR-TEM and (e) Elemental mapping of CPTHB-B2. (f) AFM and height distribution curve of CPTHB-B2.



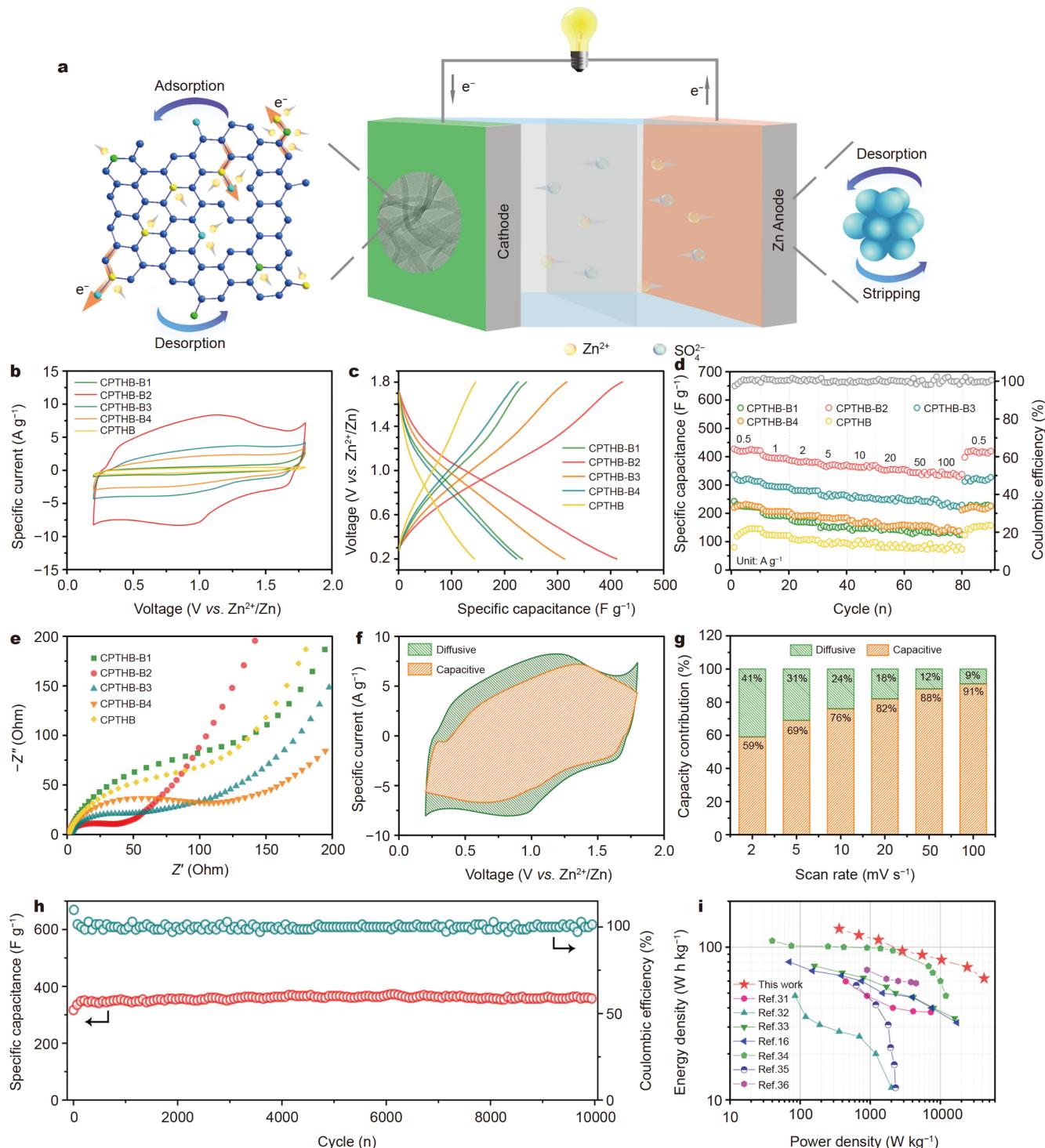
**Figure 2** (a) XRD patterns, (b) Raman spectra, (c) XPS survey spectra, (d) nitrogen adsorption-desorption isotherms, (e) BET surface areas and (f) pore size distributions of CPTHB-B<sub>x</sub> and CPTHB.

ordered carbon) and a G band (representing graphitic carbon) [25]. The peak intensity ratio of the D band to that of the G band ( $I_D/I_G$ ) is a useful indicator to evaluate the disorder of carbon materials, and the increasing value (from 0.873 to 0.910) indicates that CPTHB-B<sub>x</sub> possess low graphitic degree and high defect density caused by the B atom doping [25]. Furthermore, the chemical compositions of CPTHB and CPTHB-B<sub>x</sub> were investigated by X-ray photoelectron spectroscopy (XPS). As shown in Fig. 2c and Table S1, C (285.2 eV), O (532.2 eV), N (400.2 eV), B (192.7 eV) signals are detected in the survey spectrum, implying the successful doping of B heteroatoms [26]. The nitrogen sorption isotherms in Fig. 2d display the characteristic type I with relatively high adsorbed quantity of  $N_2$ , indicating the presence of abundant microporous structures. Besides, CPTHB-B<sub>x</sub> samples exhibit higher amount of adsorbed  $N_2$  than CPTHB. Based on the Branauer-Emmett-Teller (BET) method, BET surface areas of CPTHB, CPTHB-B1, CPTHB-B2, CPTHB-B3 and CPTHB-B4 are 791.7, 808.8, 999.2, 1105.3 and 1178.1  $m^2 g^{-1}$ , respectively, which could provide large electrochemically active surface areas for charge storage (Fig. 2e) [27]. The corresponding pore size distribution curves of CPTHB and CPTHB-B<sub>x</sub> are shown in Fig. 2f, and the porosity of CPTHB-B<sub>x</sub> is composed of the majority of micropores and few mesopores. The micropores can be regarded as the primary active sites for Zn ion adsorption and the mesopores, which can accelerate the electrolyte uptake, further increase the density of active sites [28].

As expected, the as-prepared CPTHB-B<sub>x</sub> samples exhibit an ultrathin nanobelt and heteroatom doped structure, providing sufficient Zn ion adsorption sites and rapid charge-diffusion rate. Subsequently, the electrochemical activities of the synthesized CPTHB-B<sub>x</sub> as the cathodes were evaluated in ZHSs with 2 mol  $L^{-1}$   $ZnSO_4$  as the electrolyte ( $Zn//CPTHB-B_x$ ), and the device configuration is indicated in Fig. 3a. Owing to the combination of supercapacitor-type cathode and battery-type anode, the store and output energy of ZHSs is *via* ion adsorption/des-

orption on the CPTHB-B<sub>x</sub> cathode and reversible deposition/stripping process on the Zn anode. Fig. 3b shows the cyclic voltammetry (CV) curves of CPTHB and CPTHB-B<sub>x</sub> electrodes at 20 mV  $s^{-1}$ , indicating that these ZHSs are rechargeable and able to work in a wide voltage range of 0.2–1.8 V. The current density of the CPTHB-B2 is the largest among all the samples at the same scan rate, implying that doping B atoms provides enough active sites for charge storage and an optimal doping ratio was realized. In addition, CV curves in Fig. S6 cannot display obvious deformation even at a high scan rate of 100 mV  $s^{-1}$ , which reveals the rapid kinetics and highly rechargeable feature. As presented in Fig. 3c, the specific capacitances were measured by galvanostatic charge-discharge (GCD) curves of  $Zn//CPTHB-B_x$  and  $Zn//CPTHB$  at a current density of 0.5 A  $g^{-1}$ . The specific capacitances of  $Zn//CPTHB-B1$ ,  $Zn//CPTHB-B2$ ,  $Zn//CPTHB-B3$ ,  $Zn//CPTHB-B4$  and  $Zn//CPTHB$  cathodes reach 224.8, 411.5, 313.0, 222.8 and 143.1 F  $g^{-1}$ , respectively, in agreement with the CV results. The rate capabilities of  $Zn//CPTHB-B_x$  and  $Zn//CPTHB$  were tested at different current densities, as shown in Fig. 3d (corresponding GCD curves are shown in Fig. S7). The discharge capacitances of  $Zn//CPTHB-B2$  are 411.5, 391.1, 382.2, 369.9, 364.0, 356.9, 340.7 and 320.6 F  $g^{-1}$  at current densities of 0.5, 1.0, 2.0, 5.0, 10.0, 20.0, 50.0 and 100.0 A  $g^{-1}$ , respectively, and the capacitance retention is up to 81% compared with the initial state at current densities from 0.5 to 100 A  $g^{-1}$ , which is higher than that of  $Zn//CPTHB$  (37%). When the current density returns to 0.5 A  $g^{-1}$ , a high specific capacitance of 415.3 F  $g^{-1}$  can be achieved, suggesting its high reversibility. According to the above results,  $Zn//CPTHB-B2$  demonstrates not only high capacitance, but also excellent rate performance, much better than most of reported carbon cathodes for ZHSs, which is attributed to the heteroatom doping, abundant defects and ultrathin nanobelt structures.

The electrochemical impedance spectroscopy (EIS) spectrum reveals transmission kinetics of ions and the resistance during the charging/discharging process, as shown in Fig. 3e. In the



**Figure 3** (a) Configuration and working mechanism of the as-fabricated aqueous ZHS. (b) CV curves at 20 mV s<sup>-1</sup>, (c) GCD curves at 0.5 A g<sup>-1</sup>, (d) rate performances, (e) Nyquist plots of CPTHB-Bx and CPTHB. (f) Diffusion and capacitive contribution at 10 mV s<sup>-1</sup>, (g) normalized capacity contributions at different scan rates, (h) cycling stability at 10 A g<sup>-1</sup> of CPTHB-B2. (i) Ragone plot of CPTHB-B2-based aqueous ZHS compared with other advanced aqueous ZHSs.

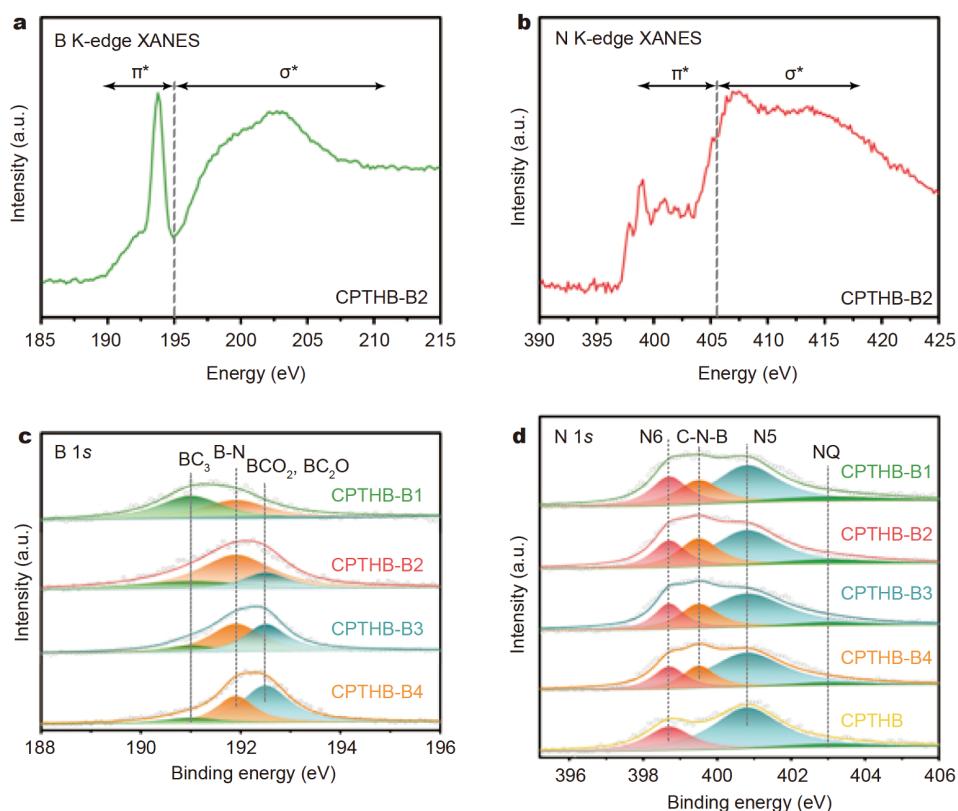
typical Nyquist plots, the semicircle diameters in the high frequency regions represent charge transfer resistance ( $R_{ct}$ ), and the slope in the low frequency regions denotes the Zn<sup>2+</sup> diffusion resistance in the electrolyte ( $R_s$ ) [29]. The  $R_{ct}$  values from the equivalent circuit model of CPTHB-B1, CPTHB-B2, CPTHB-B3, CPTHB-B4 and CPTHB based ZHSs are about 138, 34, 61, 101

and 104  $\Omega$ , respectively, demonstrating the fast charge transfer kinetics. To confirm the capacitive effect for the CPTHB-B2 cathode, quantized calculations were carried out based on the CV curves at different scan rates according to the equations in Supporting information (Equations (S4) and (S5)) [30]. As shown in Fig. 3f, g, the capacitive contribution was calculated to

be almost 76% of the total charge storage at 10 mV s<sup>-1</sup>. Besides, the capacitive contribution increases from 59% to 91% with the scan rates increasing from 2 to 100 mV s<sup>-1</sup>, suggesting the capacitive-dominated charge storage process and relatively fast kinetics under high scan rates. Fig. 3h and Fig. S8 further show the long cycling performance of the optimized ZHSs, and the Zn//CPTHB-B2 exhibits outstanding cycling stability with a high capacitance retention (98.5%) and Coulombic efficiency (~100%) after 10,000 charge/discharge cycles at 10 A g<sup>-1</sup>. Such superior capacity and rate capability of the ZHSs are very competitive, better than the majority of other reported carbon-based devices. More encouragingly, the Zn//CPTHB-B2 device provides an ultrahigh power density of 42.1 kW kg<sup>-1</sup> at 62.4 W h kg<sup>-1</sup> and a high energy density of 131.9 W h kg<sup>-1</sup> at 361 W kg<sup>-1</sup> based on the mass of CPTHB-B2. The maximum energy density and power density are better than those of Zn//CPTHB-B1 (65.1 W h kg<sup>-1</sup>, 35.0 kW kg<sup>-1</sup>), Zn//CPTHB-B3 (100.4 W h kg<sup>-1</sup>, 39.9 kW kg<sup>-1</sup>), Zn//CPTHB-B4 (69.7 W h kg<sup>-1</sup>, 32.5 kW kg<sup>-1</sup>) and Zn//CPTHB (44.7 W h kg<sup>-1</sup>, 39.3 kW kg<sup>-1</sup>) (Fig. S9). As exhibited in the Ragone plots, these values also outperform most of aqueous ZHSs and represent the best power density for all zinc-ion energy storage devices that have been reported so far (Fig. 3i and Table S2) [16,18,28,31–38]. We also compared other types of supercapacitors in Table S3, showing that ZHSs possess both high security and competitive electrochemical performance [39–43]. These inspiring results reveal that the CPTHB-B2 electrode exhibits rapid dynamics during the electrochemical process, more active sites for charge storage and high reversibility, which can be ascribed to its ultrathin ribbon structure, high heteroatom doping level, rich defects, relatively large surface area and wide pore size distribution.

To differentiate active sites for Zn ion adsorption in B, N, O co-doped carbon materials and investigate the influence of doping amount of B, we further studied the chemical states of CPTHB-Bx with different ratios of doping B through XPS and X-ray absorption near edge structure (XANES) spectra. Fig. 4a and Fig. S10a show the B K-edge XANES spectrum and peak fitting results of the CPTHB-B2, which can be divided into two regions. The sharp absorption features below 195 eV and the broad absorption peak is above 195 eV, corresponding to transition of B 1s electron to  $2p_z \pi^*$  orbitals and to  $\sigma^*$  orbitals, respectively. In the  $\pi^*$  transition domain, a sharp absorption peak at 193.8 eV is a characteristic feature of  $sp^2$  hybridized trigonal borate species (peak 2). Two additional shoulder peaks are also identified at 191.3 eV (peak 0) and 192.4 eV (peak 1), which are attributed to similar  $sp^2$  hybridized trigonal  $BC_2O$  and  $BCO_2$  structures, respectively. In the  $\sigma^*$  transition domain, the broad feature at 196–200 eV originates from trigonal B species, and the other feature at 200–205 eV is from  $sp^3$  hybridized tetrahedral B species [44–46]. This suggests the successful doping of B atoms into the carbon framework.

Similarly, N K-edge XANES of CPTHB-B2 spectrum and peak fitting results in Fig. 4b and Fig. S10b can be divided into two regions, the sharp absorption features below 404 eV with the broad absorption peak above 404 eV, corresponding to the transition of N 1s electron to  $2p_z \pi^*$  orbitals and to  $\sigma^*$  orbitals, respectively. In the  $\pi^*$  transition region, two sharp peaks at 397.8 eV (peak 1) and 398.9 eV (peak 2) are assigned to pyridinic and pyridone N species, respectively. The absorption peak at around 400.5 eV (peak 3) is generally recognized as graphitic N species, and the shoulder peak at about 402.0 eV (peak 4) is possibly associated with oxidized pyridinic N species. The last



**Figure 4** (a) B K-edge XANES spectrum and (b) N K-edge XANES spectrum of CPTHB-B2. (c) B 1s XPS of CPTHB-Bx. (d) N 1s XPS of all samples.

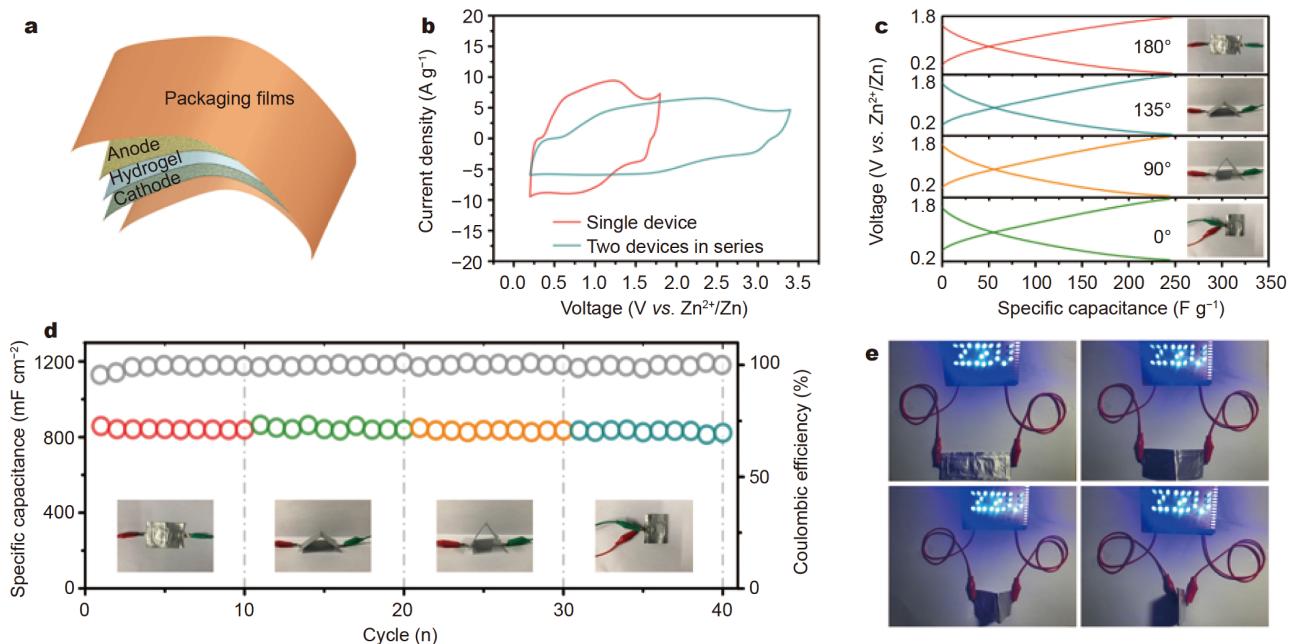
two peaks at 405.0 eV and 407.0 eV represent the transition to  $\sigma^*$  orbitals [47–49].

XPS results are displayed in Fig. 4c, d and Fig. S11. The B 1s spectra (Fig. 4c) reveal the presence of three peaks corresponding to  $\text{BC}_3$  groups (191.0 eV), B–N groups (191.9 eV) and  $\text{BCO}_2$ ,  $\text{BC}_2\text{O}$  groups (192.5 eV). The N 1s spectra (Fig. 4d) of the samples have been resolved into four individual component peaks representing pyrrolic N (N6, 398.7 eV), C–N–B (399.5 eV), pyridinic N (N5, 400.8 eV) and quaternary N (NQ, 403.0 eV). The C 1s spectra (Fig. S11a) can be deconvoluted into five peaks, which correspond to C–B groups (284.2 eV), C  $\text{sp}^2$  (284.7 eV), C–N groups (285.3 eV), C–O groups (286.3 eV) and C=O groups (288.6 eV). The spectra of O 1s (Fig. S11b) display four characteristic peaks, corresponding to O=C groups (531.4 eV), O–C groups (532.5 eV), O–B groups (533.3 eV) and absorbed oxygen (533.7 eV), respectively [26,50,51]. The results confirm that the B, N and O were incorporated into carbon frameworks, which is beneficial to favorable electrochemical performance.

For further investigation, the concentrations of B species calculated from the corresponding B 1s XPS peaks are presented in Table S4. In comparison with other B doped samples, CPTHB-B2 possesses the highest B–N content. In the meantime, more O atoms were introduced owing to the increasing mixed content of  $\text{H}_3\text{BO}_3$ , promoting the formation of more B–O groups. As mentioned above, it is found that the B–N groups are essential to enhance Zn ion adsorption and electron transport through the carbon. To verify that the active sites for Zn ion storage in CPTHB-Bx are B–N groups, two additional control experiments were carried out. The CPTHB-N-B with more N atoms and CTHB-B without N atoms were prepared and the structural characterizations are presented in Fig. S12 and Table S5. The electrochemical performances of CPTHB-N-B and CTHB-B for ZHSs were evaluated and compared with that of CPTHB-B2. As shown in Figs S13 and S14, Zn//CPTHB-N-B

exhibits a remarkable specific capacitance of  $449.7 \text{ F g}^{-1}$  at the current density of  $0.5 \text{ A g}^{-1}$  and excellent rate properties with the specific capacitance retention up to 86% when the current densities increase from  $0.5$  to  $100 \text{ A g}^{-1}$ , which is higher than that of Zn//CPTHB-B2 (81%) under the same testing condition. While the Zn//CTHB-B shows a much lower specific capacitance of  $175.2 \text{ F g}^{-1}$  and the capacitance drops to  $10.7 \text{ F g}^{-1}$  at  $20 \text{ A g}^{-1}$ . These results also verify that B–N groups are the main Zn ion hosts. To investigate the structural change on CPTHB-B2 cathode of ZHS, we characterized XRD and Raman for CPTHB-B2 cathodes at different voltages during the charge/discharge process (Fig. S15). The main XRD peaks of the CPTHB-B2 cathode remain unchanged, and the reversible emergence and disappearance of  $\text{Zn}_4\text{SO}_4(\text{OH})_6 \cdot 4\text{H}_2\text{O}$  peak is observed. Meantime, the *ex-situ* Raman spectra exhibits the same results, which demonstrates stable structure of the CPTHB-B2 cathodes during the charge storage process [38,52].

The security and flexibility of energy storage device is a demand associated with the development of portable electronic products. To meet this requirement and test the actual energy storage properties of the as-obtained CPTHB-B2 electrode, a flexible solid-state pouch cell ZHS was assembled using CPTHB-B2 as the cathode, Zn-deposited carbon cloth as the anode, and solid-state  $\text{ZnSO}_4$ /polyacrylamide (PAM) gel as the electrolyte and the separator, and then sealed by Al-plastic films under vacuum (Fig. 5a). CV curves of the device shown in Fig. 5b are similar to those of the coin cell system without distortion, indicating the regular operation of the ZHS device, and two devices connected in series can supply a voltage window of 3.4 V. Fig. 5c, d further display the electrochemical performance of the flexible solid-state Zn//CPTHB-B2 pouch cell, which presents an excellent specific capacitance of  $245.7 \text{ F g}^{-1}$  at  $5 \text{ A g}^{-1}$ . In addition, the solid-state ZHS experienced nearly no capacitance decay under different degrees of bending states. As a demonstration, two Zn//CPTHB-B2 devices connected in series



**Figure 5** (a) Structural diagram of the flexible solid-state ZHS pouch cell. (b) CV curves of two single devices in series and in parallel at the scan rate of  $20 \text{ mV s}^{-1}$ . (c) Discharge curves at  $5 \text{ A g}^{-1}$  under different deformation states, (d) cycling performance at  $5 \text{ A g}^{-1}$  under different mechanical deformations of the device. (e) Photographs of an LED pattern powered by two devices in series.

could normally power a series of light-emitting diodes (LEDs) even under bending states, verifying the practical viability of our flexible solid-state ZHSs pouch cells (Fig. 5e).

## CONCLUSIONS

In summary, we synthesized a series of ultrathin carbon nanobelts with B, N, O doping (CPTHB-Bx) as the cathodes for aqueous ZHSs. Owing to more Zn ion adsorption sites and ultrafast charge-transfer created by the heteroatom doping, the optimized CPTHB-Bx electrode displayed substantially impressive rate performance and extraordinary specific capacitance. We have successfully identified that B–N group is capable of promoting the Zn ion adsorption, and acts as the main active site for energy storage. The as-fabricated ZHS with CPTHB-B2 cathode presents an outstanding specific capacitance of 415.3 F g<sup>-1</sup> at 0.5 A g<sup>-1</sup>, and a considerable capacitance retention of 81% at the current density of 100 A g<sup>-1</sup>. In addition, a significant energy density of 131.9 W h kg<sup>-1</sup>, and a peak power density of 42.1 kW kg<sup>-1</sup> are afforded. A flexible solid-state ZHS pouch cell was assembled with the CPTHB-B2 electrode, which offers a stable cycling performance under different bending states, and two in series connected devices can power a set of LEDs. Our study can be generally applied for the design of heteroatom-doped carbon electrodes for enhancing energy storage performance in various systems.

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## 自组装杂原子掺杂碳纳米带作为超快充电锌离子混合超级电容器阴极

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**摘要** 锌离子混合超级电容器(ZHSs)由于其具有较高安全性, 较低成本和超长寿命等优点, 适合应用于大型储能设备。然而, 目前阴极材料的倍率性能较差, 严重阻碍了ZHSs的发展。基于此, 本文设计合成了一系列自组装杂原子掺杂(B,N,O)碳纳米带CPTHB-Bx, 作为ZHSs的阴极材料。杂原子掺杂可以显著改善碳骨架的化学特性, 产生更多的活性位点, 并加速电荷传输。除此之外, 本文还证明了B-N基团是快速吸附和脱附锌离子的主要活性位点。以CPTHB-B2作为阴极的ZHSs表现出最优异的电化学性能, 在0.5 A g<sup>-1</sup>电流密度下, 比电容高达415.3 F g<sup>-1</sup>, 当电流密度从0.5增大到100 A g<sup>-1</sup>时, 比电容保留率高达81%, 能量密度为131.9 W h kg<sup>-1</sup>, 功率密度为42.1 kW kg<sup>-1</sup>。此研究为超快速锌离子存储材料提供了新的设计思路。