PAPER

Effects of Crystal Size on Methanol to Hydrocarbon Conversion over Single Crystals of ZSM-5 Studied by Synchrotron Infrared Microspectroscopy†

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Operando synchrotron infrared microspectroscopy (OIMS) was used to study the conversion of methanol over coffin-shaped HZSM-5 crystals of different sizes: *large* (~ 250 × 80 × 85 µm³), *medium* (~ 160 × 60 × 60 µm³) and *small* (~ 55 × 30 × 30 µm³). The induction period, for direct alkene formation by deprotonation of surface methoxy groups, was found to decrease with decreasing crystal size and with increasing reaction temperature. Experiments with a continuous flow of dimethylether showed that evolution of the hydrocarbon pool and indirect alkene formation is also strongly dependent on crystal size. These measurements suggest that the hydrocarbon pool formation and indirect alkene generation should be almost instantaneous at reaction temperatures used in practical catalysis with crystal sizes typically ~ 1 µm³.

Key words: synchrotron, infrared microspectroscopy, ZSM-5, methanol to hydrocarbons, crystal size

1. Introduction

In-situ infrared spectroscopy is a powerful tool in zeolite catalysis, particularly for the observation of adsorbed species under reaction conditions.[1] Measurements by diffuse reflectance (DRIFTS) give spectra summed over multiple crystals contained in a loose powdered sample, while transmission measurements (TIR) examine a pressed pellet. Neither of these macroscopic sampling techniques can resolve spatial heterogeneities in the sample, and the time resolution of both is restricted by the fact that not all crystals in the sample can respond simultaneously to rapid changes in reactant concentration. Microspectroscopy, on the other hand, focusses the infrared beam onto one individual microcrystal at a time, allowing any spatial inhomogeneity between crystals or within large crystals to be resolved. Oriented groups in the framework or reacting adsorbed species can be determined via linearly polarised infrared radiation on zeolite single crystals.[2,3] Signal-to-noise limitations in conventional laboratory infrared microscopes using a blackbody (Globar) source have prevented the rapid scanning needed for studying reactions of adsorbed species in situ.

Infrared radiation from a synchrotron source is up to 100 times brighter than that from a laboratory Globar source, an improvement of a factor of ~ 20 in the signal-to-noise ratio. The first applications of synchrotron infrared microspectroscopy to zeolite single crystals were reported by Stavitzky et al.[4] This and subsequent work has exploited the *spatial* resolution of the technique down to 3-5 µm.[5–7] We have recently shown that high *temporal* resolution can also be achieved with synchrotron infrared microspectroscopy applied to the reaction of methanol in individual crystals of ZSM-5. By coupling the infrared probe with simultaneous mass spectral analysis of evolved reaction products, we could observe the initial steps in the conversion of methanol to hydrocarbons (MTH) within individual crystals on a time scale down to 0.25 s.[8]

The MTH process has been developed and commercialized over ZSM-5 for the production of gasoline due to the concerns about the shortage of oil supply during the oil crisis in the late 1970s,[9] and more recently dimethylether-to-hydrocarbons over SAPO-34 for production of light olefins with rapid industrial applications demonstrated in China.[10] The reaction mechanisms have been widely studied since.[11] There is general agreement that, under steady state working conditions, alkenes and aromatics are produced from a 'hydrocarbon pool' (HCP) within the zeolite pores, comprising a mixture of cyclic alkene and aromatic hydrocarbons from which reaction products are cracked and/or desorbed. It is less well understood how the hydrocarbon pool is formed from methanol, and in particular how the first carbon-carbon bonds are formed.[12,13] We showed that the surface methoxy groups (SMS) formed by dissociation of methanol at Brønsted acid sites at 473 K or above are the key

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intermediates in the subsequent catalysis,[8] although there is also evidence for some SMS formation at lower temperatures.[14,15] Deprotonation of the methoxy groups leads to direct alkene formation via "carbene-like" species. The alkenes can desorb or oligomerise, and the oligomer can subsequently crack or cyclise to form the cyclo-alkenyl and aromatic species which constitute the hydrocarbon pool.[8]

Our previous studies investigated the MTH reaction over HZSM-5 crystals *ca.* $160 \times 60 \times 60$ mm^3 in dimensions, prepared via synthesis in an aqueous fluoride solution. To establish the relevance of this SR IR microspectroscopic method to the investigation of reactions over zeolite crystals closer in size to those used commercially (typically <1 μm^3), a comparison was made here between HZSM-5 crystals of a range of sizes, down to ca. 30 μm in their 'middle' dimension, which approaches the limit that can be studied with the available experimental set-up (above the optical diffraction limit size of the infrared focal spot, and with overall sample thickness that guarantees an absorbance spectrum has a sufficient signal-to-noise ratio during a time resolved study). This was made possible by the development of a synthetic method that allowed the crystal size to be varied without significantly changing the framework composition (the Si/Al ratio), and thereby enabled the relative importance of reaction and diffusion processes to be varied. Furthermore, whereas the ability to prepare ZSM-5 crystals with dimensions of *ca.* 100 μm is relatively rare among high silica zeolites, <30 μm is a more accessible synthetic target.

2. Experimental

2.1 Synthesis of *large*, *medium* and *small* HZSM-5 crystals. Twinned crystals of ZSM-5 were prepared with *large* (ca. $250 \times 80 \times 85$ μm^3), *medium* (ca. $160 \times 60 \times 60$ μm^3) and *small* crystal size (ca. $55 \times 30 \times 30$ μm^3) by a modification of the ammonium fluoride route reported by Guth et al.[16] and Losch et al.[17] In a typical procedure, tetrapropylammonium bromide (TPA, 0.27 g, 1.0 mmol), NH$_4$F (0.28 g, 7.4 mmol), Al(NO$_3$)$_3$.9H$_2$O (0.19 g, 0.7 mmol) were added to deionised water (9.00 g, 500 mmol) and fumed silica (0.60 g, 10 mmol) was slowly added. The gel composition (100 SiO$_2$: 5 Al(NO$_3$)$_3$. 9H$_2$O : 75 NH$_4$F : 5000 H$_2$O : 9.1 TPABr) corresponded to Si:Al ~ 20:1. The gel was stirred at room temperature for 2 h before being sealed in a Teflon lined autoclave (30 mL, Parr) and heated at 463 K for 10 days under static conditions, yielding the medium crystal size.

A smaller crystal size was obtained by increasing the amount of Al(NO$_3$)$_3$ in the synthesis gel (Si:Al ~ 12:1), resulting in crystals of similar shape but significantly reduced volume, creating shorter diffusion pathlengths. Despite variations in the Si:Al ratio in the synthesis gels (Table 1), it was found that the crystals do not incorporate more than four Al atoms per unit cell, and any excess aluminium precipitates out as larger clumps of material which can be removed by sieving. Finally, large crystals (ca. 35× larger in volume compared to the small) were obtained from a synthesis gel with Si:Al ~ 20:1, at prolonged synthesis time (28 days at 463 K). The crystals were recovered by vacuum filtration, washed with deionised water and dried overnight at 373 K. Prior to spectroscopic studies, crystals were calcined by heating at 823 K (reached at 2 K min^{-1}) in oxygen for 64 h.

Table 1 Synthesis gel molar ratios for *large, medium* and *small* crystals

Gel comp./ Sample ID	SiO$_2$:	Al^{3+}:	F$^-$:	H$_2$O:	TPA$^+$
Large	100	5	45	2526	9.1
Medium	100	5	75	5000	9.1
Small	100	8.5	75	5000	9.1

2.2 Sample characterisation methods. Samples were characterised by thermogravimetric analysis, CHN analysis, X-ray powder diffraction (Stoe with Cu K$_\alpha$ radiation), scanning electron microscopy (JEOL JSM-5600 SME with Oxford INCA Energy 200 electron dispersive X-ray analysis) and Ar ion milling was used to cleave the crystal and study the elemental composition from the entire cross section of the crystal (Hitachi S-5500). Solid-state MAS NMR spectra were acquired on a Bruker Advance III 400 MHz spectrometer equipped with a 9.4 T widebore superconducting magnet. Prior to obtaining quantitative ^{27}Al MAS NMR, the sample was left to hydrate overnight in a moist atmosphere. The Brønsted acid sites were titrated with ammonia and quantified on a Micrometrics AutoChem 2920 coupled to a Pfeiffer Vacuum ThermoStar™ quadrupole mass spectrometer. N$_2$ adsorption isotherms were measured volumetrically (Micrometrics Tristar) after dehydrating the sample under vacuum at 473 K overnight. Catalytic tests were performed on a Salamander fixed-bed microreactor. The calcined crystals were diluted with quartz and tested for the dimethyl-ether-to-hydrocarbons reaction at weight hourly space velocity (WHSV) of 3.8 h^{-1}. Gas phase products were separated and analysed by on-line gas chromatograph (Agilent GC) equipped with a capillary GC GasPro column (0.32 mm × 30 m) using a Flame Ionisation Detector (FID).

2.3 *Operando* Infrared Microspectroscopy (OIMS) experiment. HZSM-5 crystals were sprinkled on CaF$_2$ windows and placed on the heated sample stage of an environmental Linkam FTIR600 cell (c.a. 50 mL internal volume) mounted on a (Bruker) Hyperion 3000 IR microscope coupled to a Vertex 80V FTIR at MIRIAM beamline B22 of Diamond Light Source. The infrared experiment requires only single crystals of the catalyst. To improve the sensitivity of the MS detection we typically added around 1 mg (up to 3 mg) of zeolite to the cell and checked that the spectra seen from one crystal were reproducibly seen from other crystals in the sample. The infrared experiments were collected in transmission from individual crystals at spatial resolution down to 10×10 micron at the sample using a 36× objective/condenser optics, averaging 16 scans at 4 cm^{-1} resolution, i.e. typically collecting an IR spectrum every two seconds. The output of the gas phase products from the cell was analysed by mass spectrometry (Pfeiffer or EcoSys) and correlated with the time-resolved IR spectral changes. Gases or liquids were introduced into a nitrogen flow (100 mL min^{-1}) upstream of the heated cell at temperatures up to 673 K and ambient pressure. Multiple methanol pulses (each pulse 8 μL in volume) were syringe injected, and dimethylether (5 mL min^{-1}) was diluted in the nitrogen stream and introduced in a continuous flow.

3. Results

3.1 Structure, calcination and unit cell composition.

Upon removal of the structure directing agent used in the synthesis (Fig. S1 TGA and Table S1 CHN analysis, ESI†), the protonic form of ZSM-5 (here HZSM-5) is obtained with a high surface area (~350 $m^2 g^{-1}$) as confirmed by nitrogen adsorption (Fig. S2 for XRD and Fig. S3 for N_2 adsorption, ESI†). SEM micrographs of the three crystals display a highly uniform crystal size distribution for each crystal batch (Fig. 1). This is important as a separate sample of freshly-calcined HZSM-5 crystals was used for each IR experiment.
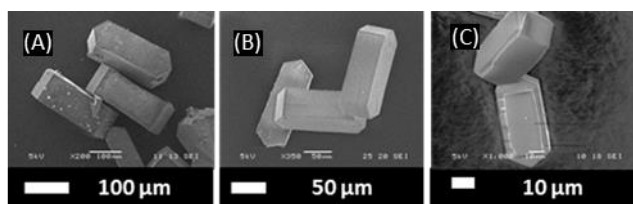
Fig. 1 SEM images of (A) large, (B) medium and (C) small HZSM-5 crystals.

The ^{27}Al and ^{29}Si NMR spectra of the crystals are shown in Fig. 2. The [27]Al peak at 53-56 ppm confirms Al is tetrahedrally coordinated into the lattice in all three sets of crystals.[18,19] The [29]Si spectra of all three show the profile expected for crystalline HZSM-5 with Si: Al of ~ 25.[20] The as-prepared crystals display a classic hour-glass outline in optical micrographs (see inset in Fig. 2). This is due to intergrowths of two different orientations of the *a* and *b* axes of the crystal.[21]
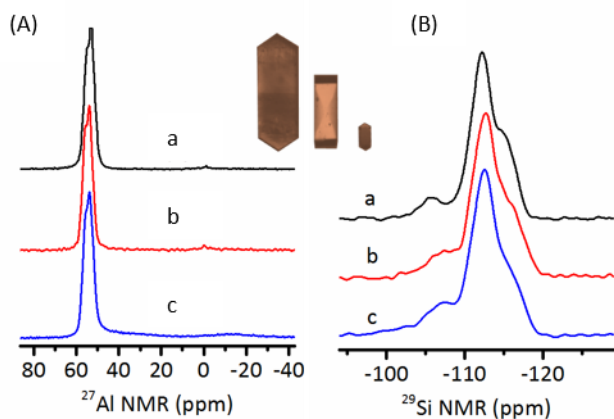

Fig. 2 ^{27}Al MAS NMR (A) and ^{29}Si MAS NMR (B) of the (a) *large*, (b) *medium* and (c) *small* HZSM-5 crystals. Inset. Optical images of the three crystals as seen under polarised light.

Elemental analysis of the crystals by bulk EDX and cross-sectional EDX on cleaved crystals gave Si: Al ratios between 25 and 36 (Table 2). Brønsted acid site concentrations determined from ammonia TPD measurements (Table S2, ESI†) gave similar Si: Al ratios, and Al line cross-sectional profiles revealed that the aluminium is homogeneously distributed across the cleaved crystals (Fig. S4, ESI†), consistent with the reported uniformity of composition achieved with the use of fluoride in the synthesis.[17]

Table 2 Characterisation of Si: Al ratio from different techniques

a × b × c (μm³)	250 × 80 × 85 *Large*	160 × 60 × 60 *Medium*	55 × 30 × 30 *Small*
EDS	34	31	30
CS-EDS	36	24	24
NH₃-TPD	24	21	21

3.2 Microcatalytic test.

The performance of the crystals in a conventional microcatalytic test was confirmed by carrying out DME conversion at 673 K. Table 3 shows conversion and product selectivity data after 5 minutes on stream (conversions decreased at longer reaction times due to deactivation). Olefins are the major products under these conditions, but significant amounts of aromatics are also formed.

Table 3 Microreactor selectivity and conversion data from dimethylether (DME) conversion over the three different crystal sizes of HZSM-5 at 5 min time-on-stream, 673 K, WHSV = 3.8 h^{-1}

Microreactor selectivity	Large	Medium (%)	Small
ethene	23.1	24.5	32.2
propene	42.9	46.0	42.9
C4[a]	20.3	19.5	18.0
C5[a]	5.6	4.9	4.7
C6+	6.3	5.1	2.15
C_2/C_3[b]	0.5	0.5	0.8
DME Conversion	99%	95%	100%

IR spectra of individual dehydrated HZSM-5 crystals. IR spectra were recorded for the *large* HZSM-5 (*ca.* 250 × 80 × 85 μm³), *medium* (*ca.*160 × 60 × 60 μm³) and *small* crystal sizes (*ca.*55 × 30 × 30 μm³) after dehydration at 623 K in flowing nitrogen. Fig. 3 compares spectra of the three dehydrated crystals. Each crystal shows two hydroxyl bands in the high wavenumber region, the sharp band at 3600 cm^{-1} is due to the Brønsted acid site (BAS) and the small band at 3735 cm^{-1} is due to terminal silanol functional groups located at the interfaces between the crystal intergrowths or other defects. Closely similar spectra were measured from different crystals within each batch, and from different regions of individual crystals. In particular, we could not distinguish any differences in the relative populations of silanol and BAS hydroxyl groups between the different areas of the intergrown crystals. (Fig. S4 for FTIR spectra of different crystals, ESI†). The lower wavenumber bands at 1995, 1870 and 1630 cm^{-1} are due to overtones and combination bands of Si-O-Al stretching and bending modes of the zeolite framework. The intensities of these bands are determined by the thickness of the zeolite crystals; the spectra in Fig. 3 have been normalised against the intensity of the 1870 cm^{-1} band, and a rubber-band base-line correction applied. Typically, spectra of the dehydrated zeolites are subtracted from those measured during reaction, and the difference spectra presented.
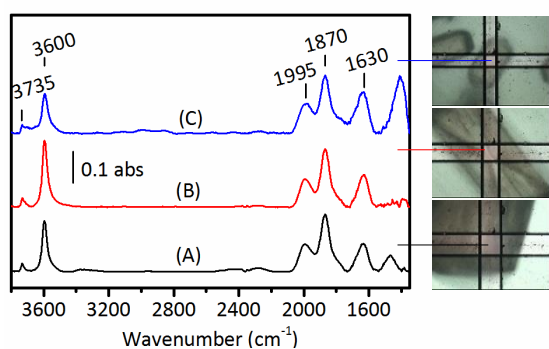
Fig. 3 SR-IR spectra of dehydrated individual HZSM-5 crystals (A) *large*, (B) *medium* and (C) *small*. Spectra are scaled based on the framework band at 1870 cm⁻¹ for comparison of the intensity of the Brønsted acid site (scaling factors: circa x3, x2 and 1 for the *large*, *medium* and *small* crystal sizes, respectively).

3.3 Methanol pulses. Fig. 4 provides an overview of the species that are detected by infrared microspectroscopy in the early stages of the MTH reaction over HZSM-5, namely adsorbed methanol, at 523 K and hydrogen bonded DME at 573 K over a *large* crystal, and methoxy groups and alkene oligomers at 573 K over a *medium* crystal. (A full sequence of spectra at 2 second intervals following injection of a methanol pulse at 523 K on a *medium* crystal is given in Fig S6, ESI†). When methanol is first injected into the zeolite it is hydrogen bonded to the BAS. At 523 K, the hydrogen bonded methanol is quickly converted to hydrogen bonded dimethylether, which slowly desorbs from the zeolite, recovering part of the original BAS hydroxyls intensity and leaving a weak residual spectrum of surface methoxy groups (SMS). After a further delay, the BAS hydroxyl band intensity dramatically increases, and the SMS spectrum is converted to that of alkene oligomer, identical to that formed when propene is injected into the zeolite at the same temperature (523 K).[8]
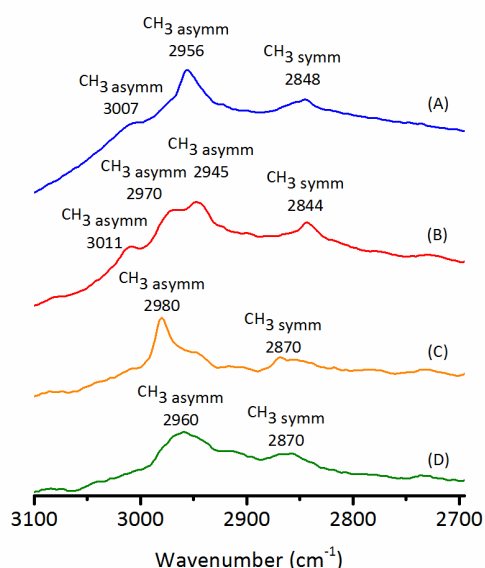
Having described the type of adsorbed species formed in the early stages of the MTH reaction over HZSM-5, next, the response of three different sizes of crystal to a pulse of methanol at 573 K is compared (Fig. 5). The top panels show the MS analyses of the effluent gases, the central panels show the intensity of the BAS hydroxyl band ~ 3600 cm⁻¹ as a function of time (expressed as a percentage of the original intensity in the dehydrated crystal) and the bottom panels show IR spectra in the CH stretching region at the reaction times indicated. In all crystal sizes studied here, oligomeric hydrocarbon species are observed immediately after the loss of methoxy groups at the end of the induction period.
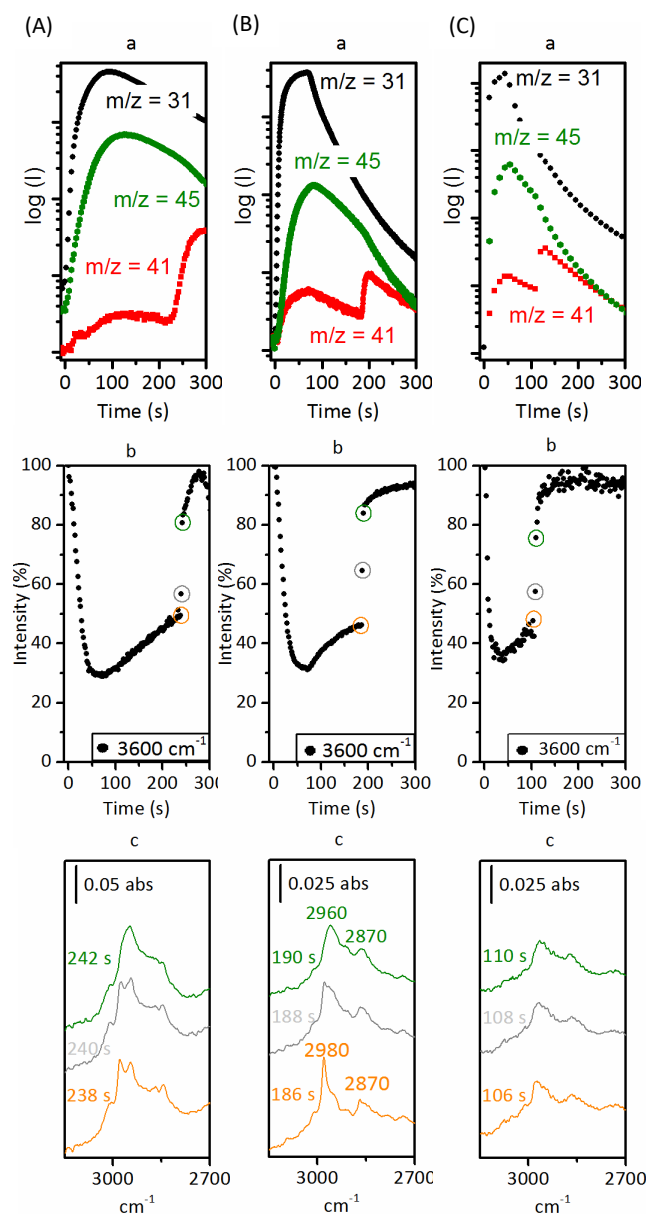

Fig. 4 Difference FTIR spectra of a *medium* size HZSM-5 crystal, showing the sequence observed following introduction of methanol at 523 K. The marked bands at each stage of the reaction are identified with (A) hydrogen bonded methanol at 523 K, (B) hydrogen bonded dimethylether at 573 K, (C) surface methoxy groups at 573 K and (D) alkene oligomers at 573 K.

Fig. 5 Comparison of (A) *large*, (B) *medium* and (C) *small* crystals during one 8 μL methanol pulse injected into a N₂ flow of 100 mL min⁻¹ at 573 K. (a) Mass spectral traces recorded during these experiments; m/z = 31 measures methanol, m/z = 45 DME, m/z = 41 propene (with a contribution from DME fragmentation). (b) Corresponding time course of the zeolite hydroxyl ν(OH) 3600 cm⁻¹ band intensity relative to an activated crystal recorded at 2 s intervals. (c) Evolution of the CH stretching bands during the rapid regeneration of the zeolite hydroxyls during the experiments in (A, B and C), respectively.

In the *large* crystal (Fig. 5 A), the loss of OH intensity following injection of methanol is followed by a slow recovery as hydrogen bonded DME is desorbed, then after an induction period of ~ 240 s there is a remarkably rapid recovery of OH intensity. The mass spectral traces during this experiment show a clear delay in evolution of DME after the methanol injection, but the first appearance of propene occurs at the same point as the recovery of OH intensity. The infrared spectrum in the ν(CH) region shows the presence of SMS and some residual hydrogen bonded DME (orange trace in Fig. 5A(c)). At 242 s all of the SMS and DME have gone and the spectrum is that of the oligomer formed from alkenes (green trace in Fig. 5A(c) and Fig. 4).

In the large crystal, the loss of SMS and residual DME occurs at the same time as the BAS hydroxyl groups are recovered and propene is evolved. This could be taken as evidence of a reaction between SMS and residual DME regenerating BAS described by eqn. (1):

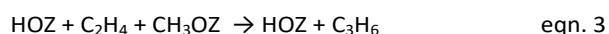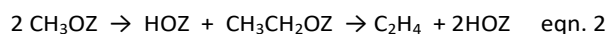$$ZOCH_3 + CH_3OCH_3 \rightarrow ZOH + CH_3CH_2OCH_3 \qquad \text{eqn. 1}$$
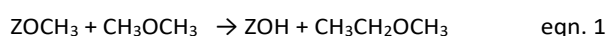
However, the corresponding data for medium and small crystals do not support this explanation (Fig. 5 B-C). In the medium crystals, the spectrum remaining at the point where the BAS OH band intensity begins to recover (orange trace in Fig. B(c) shows only SMS with little remaining DME adsorbed on the crystal is seen by IR), and this is equally clear for the small crystals.

The effect of crystal size on the induction time for propene formation, BAS hydroxyl recovery and SMS conversion to oligomer is dramatic. We attribute these differences to the different length of diffusion pathways in the different sized crystals. From these experiments the deprotonation of SMS, which initiates propene formation, appears to occur only after desorption of most or all of the initially hydrogen bonded DME, which occurs more rapidly from the medium and small crystals than from the large crystals.

Previous evidence for direct formation of hydrocarbons from SMS in ZSM-5 has come from NMR experiments of Wang et al.[22], who suggest that carbene or ylide intermediates may be involved. We note also that infrared evidence for a carbene-like reactivity of SMS in ZSM-5 towards light alkenes (e.g. methylation of ethene to form propene) has been presented by Yang et al.[23]

The direct formation of alkenes through a carbene-like intermediate requires mobility of SMS within the zeolite such as described by eqn. (2) and (3) below:

$$2\ CH_3OZ \rightarrow HOZ + CH_3CH_2OZ \rightarrow C_2H_4 + 2HOZ \quad \text{eqn. 2}$$

$$HOZ + C_2H_4 + CH_3OZ \rightarrow HOZ + C_3H_6 \qquad \text{eqn. 3}$$

The alkenes formed can desorb from the zeolite or oligomerise at BAS in the zeolite (giving the observed ν(CH) bands identical to those formed on injecting propene or ethene into the zeolite).[8] We emphasize that this chemistry is occurring only during the initial interaction of methanol with a dehydrated zeolite crystal.

Comparison of the reactivity of methanol injected into a *large* and *small* crystal at 573 K is made in Fig. 6. In both crystals, a clear induction period for propene formation is seen only in the first methanol pulse. More propene is produced in the second, third and fourth pulses, while the DME yield reduces, see MS plots in (a). The higher propene and ethene yields seen in second and subsequent pulses in Fig. 6(a) suggest that alkene formation through this indirect process is more productive than that occurring during the initial deprotonation of SMS. There is also some indication of ethene formation from the second pulse onwards (the high baseline at m/z = 27 is due to overlap with the m/z = 28 peak of the nitrogen carrier gas). Methyl aromatic products are also seen to be rising with successive methanol pulses (the m/z = 91 peak is due to the tropylium cation, which is an abundant ion rearrangement product in the mass spectra of methyl aromatic compounds).

The behaviour of the BAS OH band at 3600 cm^{-1} is also different in the second and subsequent methanol pulses from that seen in the first pulse, plots in Fig. 6(b). The recovery of BAS after each methanol pulse gradually diminishes and the number of residual BAS remaining in the zeolite declines, especially in the *large* crystal.

Two new bands appear and grow in the 1650-1450 cm^{-1} region, during the four successive methanol pulses at 573 K, plots in Fig. 6(c). The time dependence of the 1510 cm^{-1} and 1620 cm^{-1} bands is clearly different. In both crystal sizes, the 1510 cm^{-1} band appears first, and decays less rapidly than the 1620 cm^{-1} band. However, in the smaller crystal, the growth of the 1510 cm^{-1} band occurs notably faster in the first pulse. Nonetheless, the build-up of these species correlates with the declining concentration of BAS remaining after each methanol pulse, supporting the assignment that they are associated with the hydrocarbon pool (HCP) accumulating in the zeolite.

The band at 1510 cm^{-1} (with a shoulder at 1460 cm^{-1} and accompanying ν(CH) bands at 3120 cm^{-1} and 2910 cm^{-1} was assigned to the 1,3-dimethylcyclopentenyl cations (Scheme 1).[8]
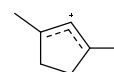
Scheme 1. Structure of 1,3-dimethylcyclopentenyl cations (DMCP).

DMCP has been identified by NMR as the first cyclic component of the hydrocarbon pool to be formed in HZSM-5 during the MTH reaction,[24] and the vibrational frequencies observed by IR microcrystal spectroscopy are close to those of the gas phase cations reported elsewhere.[25] Once DMCP forms in HZSM-5 during the MTH reaction, it may undergo methylation and skeletal rearrangement to form toluene or other methylated aromatic products. DMCP may also crack to form alkenes.[24] The 1620 cm^{-1} band was assigned to a C=C stretching mode of adsorbed methyl aromatic species.[8]

At the low temperatures investigated so far (<623 K), hydrogen bonded species were seen during the leading edge of the methanol pulse.
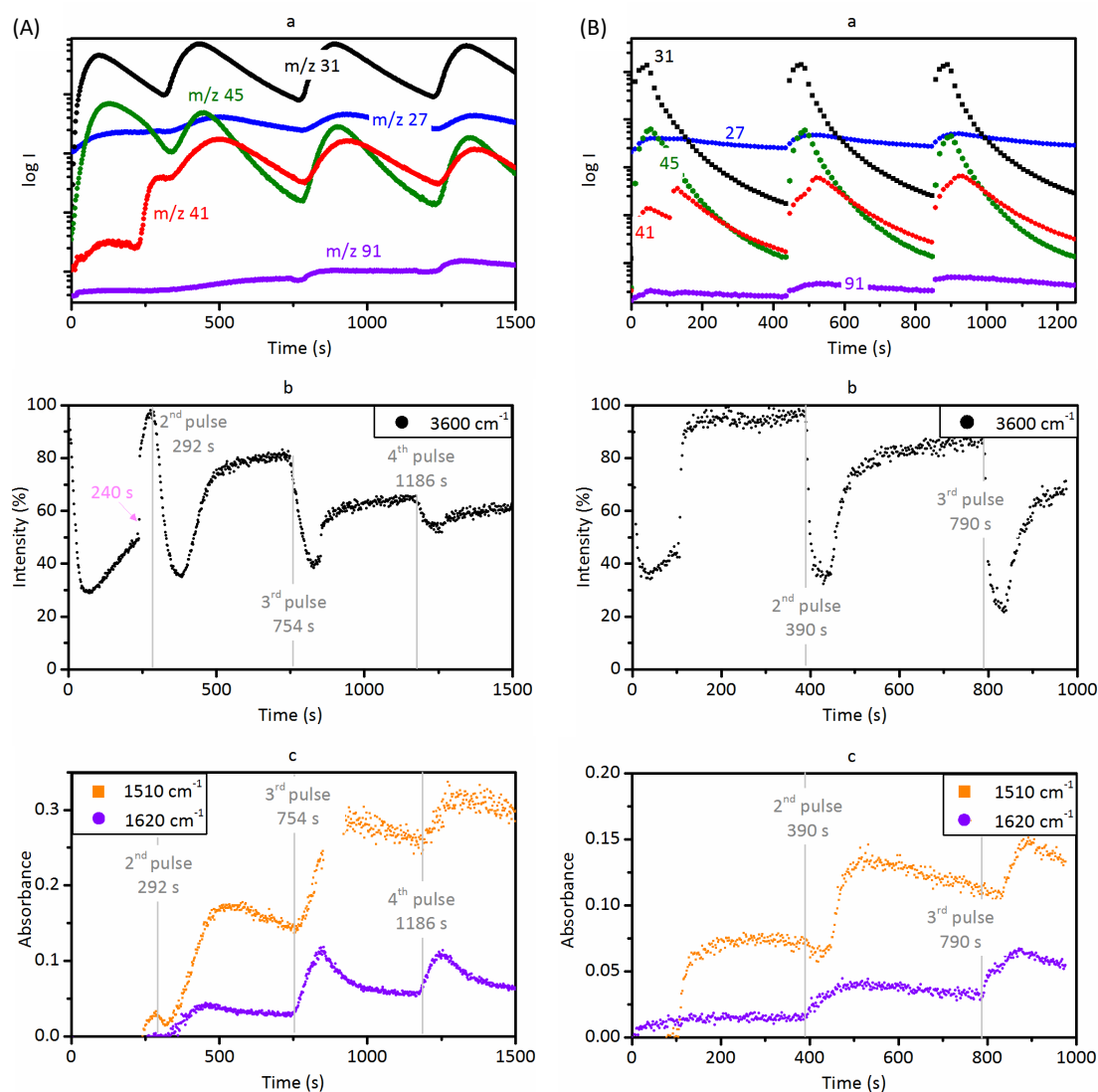
Fig 6 Comparison of a (A) *large* and (B) *small* crystal during 4 and 3 successive 8 μL methanol pulses at 573 K, respectively. (a) Gas phase products detected by online-MS during injection of methanol. (b) Time dependence of the infrared OH stretching band at 3600 cm^{-1} recorded at 2 s. (c) Extracted intensity of the bands at 1510 cm^{-1} (orange) and 1620 cm^{-1} (purple) recorded at 2 s intervals during the experiments in (A) and (B), respectively.

Considering the formation of aromatics is promoted at higher temperatures, it was desirable to compare the different crystals at higher reaction temperatures, where no hydrogen bonding is seen and the differences between the production of aromatics becomes more apparent. It was found that at higher reaction temperatures (623 K) and in smaller crystals the 1620 cm^{-1} band is more intense, and the relative yields of aromatic products at m/z = 91 detected in the MS are higher (Fig. 7). The induction period for recovery of BAS, loss of SMS and evolution of propene during the first methanol pulse at this temperature is less than 10 s. The relative yields of alkene products remain similar between the *medium* and *small* crystals, as are the 1510 cm^{-1} band intensities. However, the 1620 cm^{-1} band has two orders of magnitude (integrated area) higher relative intensities compared to the *large* and *medium* crystals, and an enhanced yield of methylaromatic products is detected in the MS for the *small* crystal. At this temperature both the 1510 and 1620 cm^{-1} bands are diminished equally rapidly, and the loss of

these bands appears to correlate with recovery of the BAS during the second and third pulses.

At low temperatures (< 523 K) methanol is converted slowly to DME, which involves the formation of SMS, but the deprotonation of the SMS occurs only after hydrogen bonded DME is desorbed. At higher temperatures hydrogen bonded DME is quickly desorbed, the formation of alkenes and oligomer follows, and in subsequent pulses of methanol the cyclic HCP species are seen.

The important conclusion regarding the significance of crystal size is that the induction time for deprotonation of SMS, regeneration of BAS and formation of alkenes decreases as the crystal size is reduced. We attribute this to the faster desorption of hydrogen bonded DME out of the smaller crystals i.e. the presence of excess DME in the crystals appears to block the SMS deprotonation chemistry.
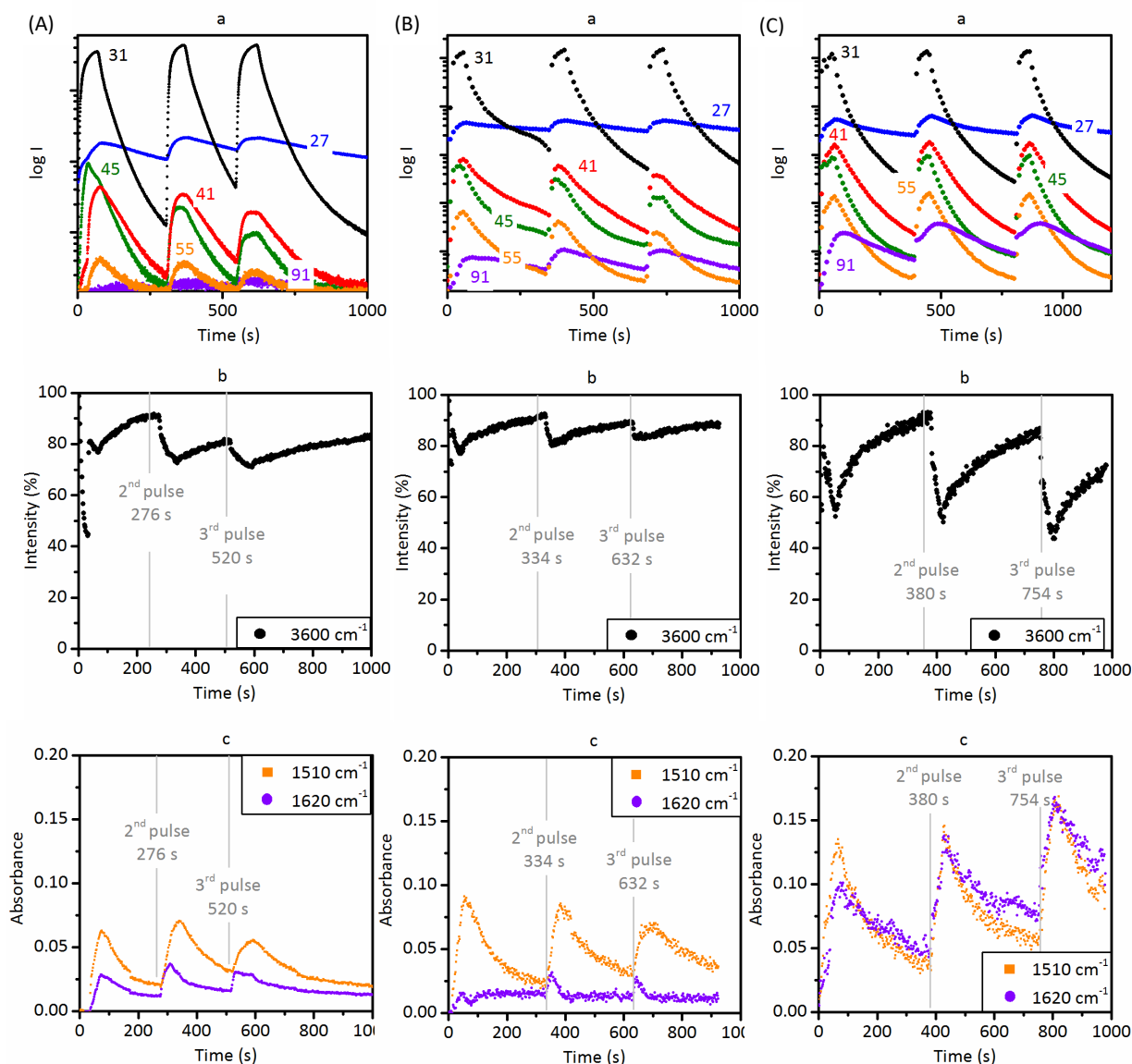
Fig. 7 Comparison of (A) *large* (B) *medium* and (C) *small* crystal of HZSM-5 crystal during the injection of 3 successive 8 µLpulses of methanol at 623 K. (a) On-line MS analysis of gas phase products. (b) BAS band intensities relative to the dehydrated zeolite. (c) Absolute band intensities of the 1510 and 1620 cm^{-1} bands recorded at 2 s intervals during the experiments in (A, B and C), respectively.

This *direct* formation of alkenes can be clearly seen in all three crystal sizes at lower temperatures but at 623 K, the typical temperature for MTH catalysis, the recovery of BAS due to deprotonation of SMS can be seen only in the initial methanol pulse in large crystals. The formation of alkenes (and aromatics) at this temperature is associated with the presence of the HCP species i.e. is *indirect.*

3.4 Continuous Flow of DME. To understand better the chemistry of the HCP species responsible for the 1510 and 1620 cm^{-1} bands and the effects of crystal size on the formation and reaction of these species under steady-state conditions, we turned to a continuous flow of DME (rather than methanol

pulses). DME is readily converted to hydrocarbons over ZSM-5 catalysts.[26] Analysis of the v(CH) region of spectra measured in the presence of DME is more difficult because of the overlapping bands of both gas phase DME and hydrogen bonded DME, and we will present a detailed investigation of these spectra elsewhere. It is useful however to examine the 1650-1450 cm^{-1} region together with the MS analyses in the context of the present discussion of crystal size effects on HCP formation (Figures 8-10).
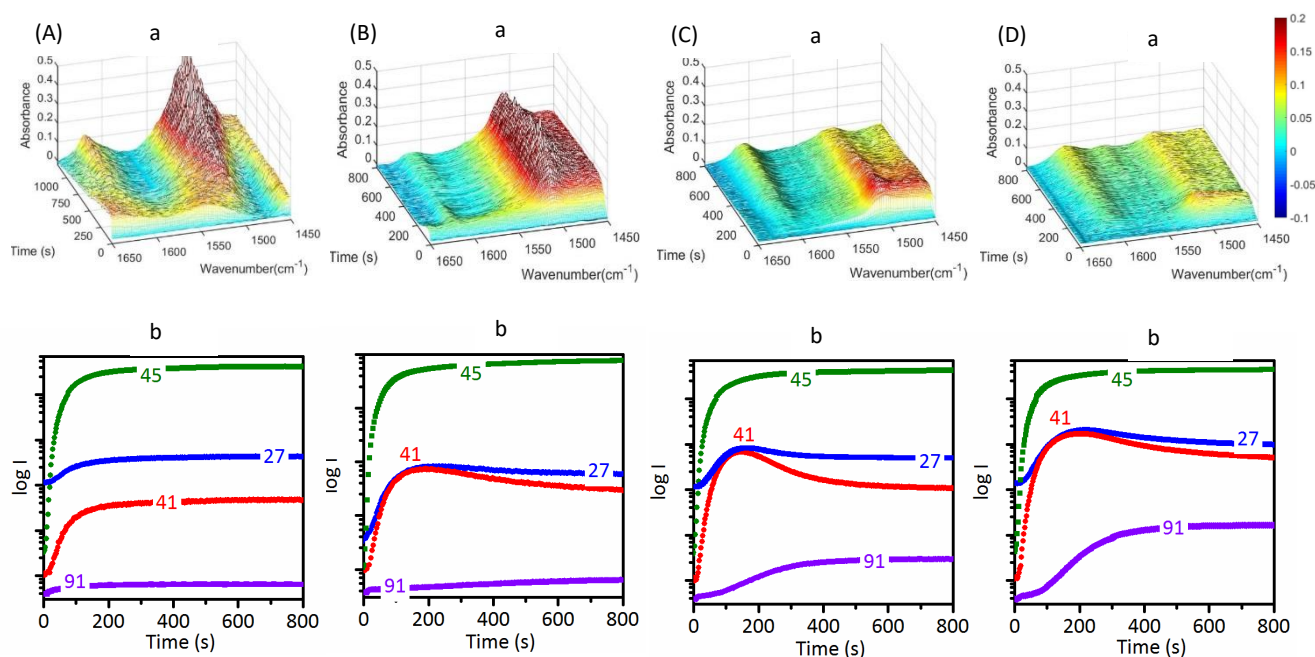
Fig 8 Continuous flow of dimethylether over a *large* crystal of HZSM-5 at (A) 523 K; (B) 573 K; (C) 623 K and (D) 673 K. (a) IR spectra recorded at 2 s intervals. (b) Online-MS analysis of gas phase products during the experiments in (A-D), respectively.
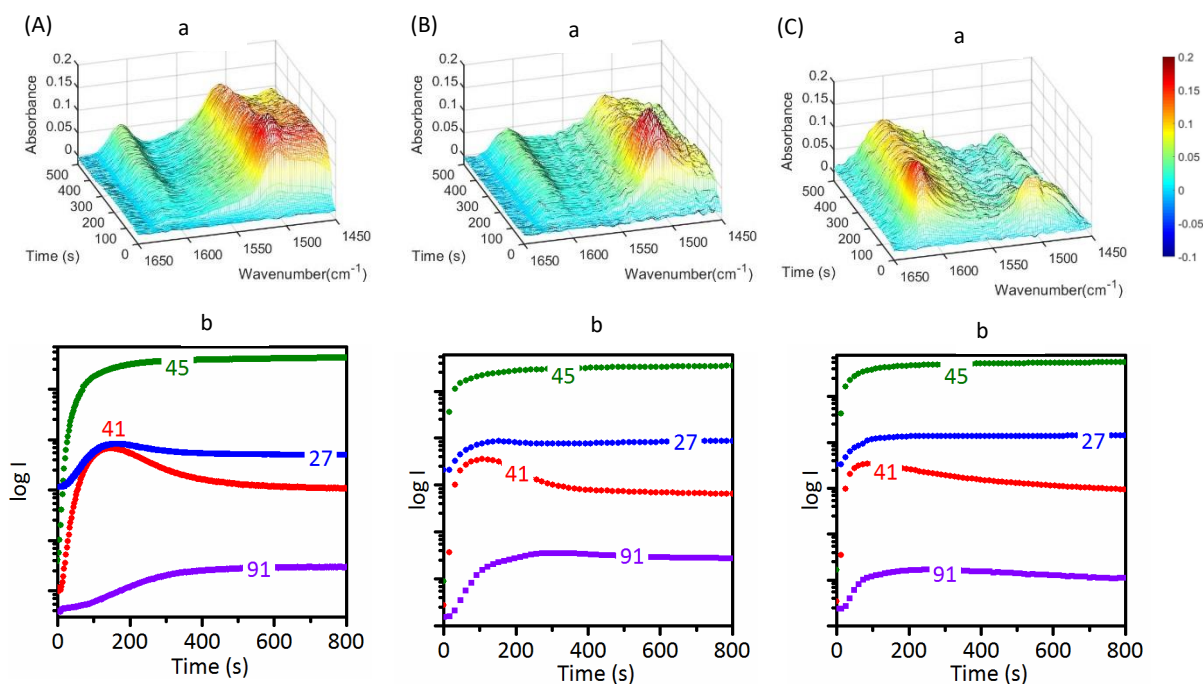

Fig. 9. Comparison of (A) *large*, (B) *medium* and (C) *small* HZSM-5 crystal during a continuous flow of dimethylether at 623 K for 8 min. (a) IR spectra recorded at 2 s intervals. (b) On-line MS analysis of gas phase products during the experiments in (A, B and C), respectively

At 523 K the 1510 and 1480 cm^{-1} bands of DMCP appear after a short delay and continue to grow in intensity over the duration of the experiment (Fig. 8-A). The MS analysis shows little evidence of alkene formation at this temperature; the observed m/z = 41 signal results from fragmentation of DME in the mass spectrometer. At this temperature, any alkenes formed are evidently oligomerised in the zeolite and slowly react to form DMCP. Also evident at 523 K is the very slow

growth of a 1620 cm^{-1} band, indicating that there is some further reaction to form aromatics at this temperature, although none escape from the zeolite. At 573 K, the 1510 and 1460 cm^{-1} bands of DMCP rise quickly, plateau and remain constant thereafter (Fig. 8-B). Weak, but evident is the 1620 cm^{-1} band of methyl aromatics which broadens and shifts to lower frequency at 1607 cm^{-1} after about 400 s of flowing DME over the large HZSM-5 crystal, suggesting further

evolution of the adsorbed methyl aromatics. The MS traces at 573 K show enhanced levels of propene and ethene, but still very little evolved methyl aromatics.

At 623 K, the steady state concentration of DMCP in the zeolite is considerably less than at lower temperatures (Fig. 8-C). There is an increased contribution to the spectrum from the 1620 cm⁻¹ band, and in the MS traces the initial yields of propene and ethene decline as the evolution of methylaromatics increases. The generation of methylaromatic species (m/z = 91) is more prominent at 673 K, but the IR spectra are more complex. Within 200 s, multiple bands are observed at 1607, 1566, 1510, 1490, 1460 cm⁻¹, suggesting formation of a more varied HCP at this temperature (Fig. 8-D).

A striking difference in the DMCP concentration between *large* and *small* crystals was observed during the isothermal continuous flow of DME at 623 K (Fig. 9). The density and strength of acid sites has not changed, and the product distributions at longer reaction times are quite similar, although the steady state evolution of methylaromatics is achieved more quickly in the smaller crystals. The decomposition of DMCP species to generate methylaromatics evidently occurs more quickly in the smallest crystals.

Similar effects were seen at 673 K (Fig. 10). In both *large* and *small* crystals a complex set of bands is seen in the 1650-1450 cm⁻¹ region but there is a larger contribution from bands in the 1550-1650 cm⁻¹ region in the *small* crystals, consistent with the more rapid evolution of methylaromatic products in the MS traces. At this temperature the crystals contain many of the HCP species detected by NMR and UV-VIS spectroscopy.[12,13]
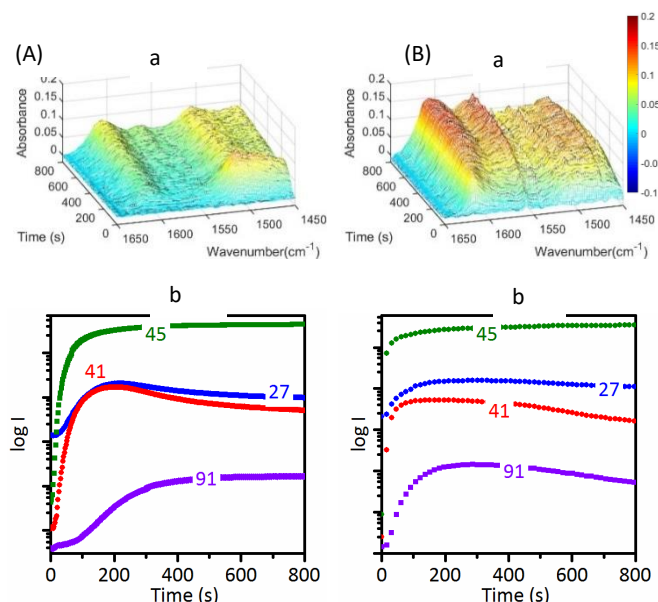

Fig 10 Comparison of (A) *large* and (B) *small* HZSM-5 crystal during continuous flow of dimethylether at 673 K over 13 min. (a) IR spectra recorded at 2 s intervals. (b) On-line MS analysis of gas phase products during the experiments in (A) and (B), respectively.

The continuous flow DME experiments show that the DMCP species which we identify as the first cyclic component of the HCP is formed in high concentrations in the zeolite already at

523 K, but with little formation of gas phase products at this temperature.

The infrared signature of adsorbed methyl aromatics (1620 cm⁻¹) begins to appear after the DMCP (1510 cm⁻¹). At higher temperatures cracking of the DMCP to form alkene products is accompanied by the further reaction to generate methyl aromatics, and both of these processes depend on crystal size. In particular, the conversion of DMCP species to methyl aromatics is favoured in smaller crystals, probably because of faster diffusion of DME reactant into the pores.

4. Conclusions

Operando infrared microspectroscopy (OIMS) performed with synchrotron radiation is a powerful tool to study the formation mechanism of the initial hydrocarbon pool species within individual crystals of ZSM-5. In this study we have confirmed the reaction sequence proposed elsewhere,[8] in which the *direct* formation of alkenes in the very initial stages of methanol conversion is due to deprotonation of SMS. Once alkenes are formed, these alkenes oligomerize at BAS; the oligomer can either crack or can cyclize to form DMCP as the first *cyclic* component of the hydrocarbon pool which can also release alkenes or react further to form the methylaromatic components of the hydrocarbon pool. We show that the OIMS enabled us to follow reactions on crystals down to 30 microns in size, with a time resolution of 2 s routinely, and down to 0.25 s.[8]

The important conclusion regarding the significance of crystal size is that the induction time for deprotonation of SMS, regeneration of BAS and formation of alkenes decreases as the crystal size is reduced. We attribute this to the faster desorption of hydrogen bonded DME out of the smaller crystals i.e. the presence of excess DME in the crystals appears to block the SMS deprotonation chemistry.

These studies have shown that it will be difficult to detect *direct* formation of alkenes in much smaller micron sized crystals, since the HCP species are seen to form very quickly in smaller crystals, and the cracking of these is responsible for the observed steady-state alkene yields from both methanol and dimethylether. Based on this, we suggest that the induction periods for alkene formation seen in many conventional catalytic studies cannot be due to intra-crystalline chemistry but must be attributed to diffusion of reactants into the catalyst bed or of products out of the catalyst bed.

Finally we note the potential of the method to be applied to other zeolite catalysts and catalytic reactions, given crystal sizes of at least 20 microns for satisfactory signal-to-noise ratios, and the possibility of extending the measurements to time dependent spectra from several adjacent crystals using automatic mapping functionality common to the IR microscope acquisition software. The signal to noise advantage gained from synchrotron radiation will also allow diffusion measurements on zeolite crystals previously performed with laboratory infrared microscopes[27,28] to be extended to much shorter time scales.

ARTICLE Journal Name

Acknowledgements

There are no conflicts to declare.

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